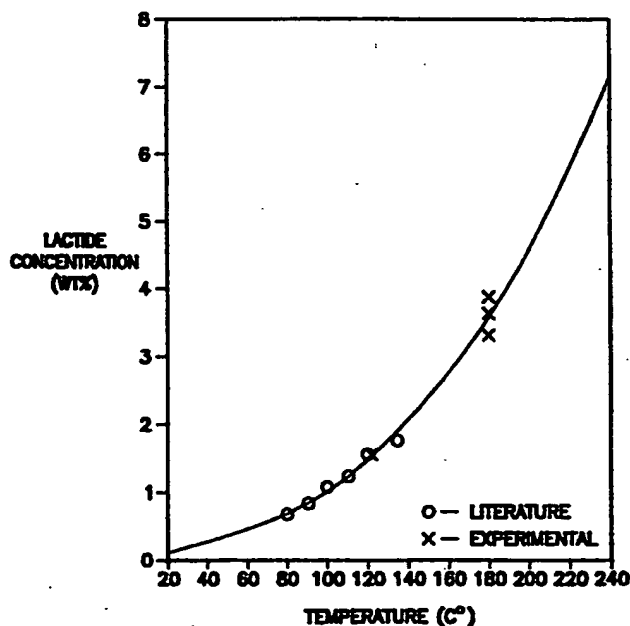




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(54) Title: MELT-STABLE LACTIDE POLYMER FILMS AND PROCESSES FOR MANUFACTURE THEREOF



(57) Abstract

Amorphous and semi-crystalline films comprised of a lactide polymer. The lactide polymer comprises poly(lactide) polymer chains, residual lactide in concentration of less than about 5 percent by weight and water in a concentration not greater than about 2000 parts-per-million. Processes for manufacturing amorphous and semi-crystalline films with the lactide polymer composition are also disclosed.

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**MELT-STABLE LACTIDE POLYMER FILMS AND PROCESSES
FOR MANUFACTURE THEREOF**

BACKGROUND OF THE INVENTION

5 1. Field of the Invention

The present invention relates to films comprising a melt-stable, biodegradable, lactide polymer composition and processes for manufacturing the films from the melt-stable, biodegradable polymer.

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2. Description of the Prior Art

The need for polymeric biodegradable films is well established. Films manufactured from blown or cast processes are well known. Typically in a blown film process, a plastic melt passes through a die which extrudes the molten plastic into an annular shape. Typically, the extruded film is extruded in an upward fashion. As the film moves upward, air is blown into the film which expands the film into a tubular shape. The tube is generally closed at some distance above the die, with a pair of nip rolls.

In a cast film process, a sheet is typically extruded from a slit die. The sheet is thereafter pulled through a series of rollers which cool the extruded sheet and may also elongate the length and width of the sheet to a desired dimension and thickness.

The use of films is widespread and well known in the art. The heaviest use of films occurs in the packaging and disposable article industries. Films employed in the packaging industry include films used in food and non-food packaging, merchandise bags and trash bags. In the disposable article industry, the general uses of films occur in the construction of diapers and personal hygiene articles, including tapes.

35 In light of depleting landfill space and adequate disposal sites, there is a need for biodegradable films.

Currently, films comprising polymers such as polyethylene, polypropylene, polyethylene terephthalate, nylon, polystyrene, polyvinyl chloride and

polyvinylidene chloride are popular for their superior extrusion and film-making properties. However, these films are not biodegradable. Furthermore, these films are generally noncompostable, which is undesirable from an environmental point of view.

Films have been developed which are generally considered to be biodegradable. These are films which purportedly have adequate properties to permit them to break down when exposed to conditions which lead to composting. Examples of such arguably biodegradable films include those made from polycaprolactone, starch biopolymers and polyvinyl alcohol.

Although films extruded from these materials have been employed in film containing articles, many problems have been encountered with their use. Often the films are not completely biodegradable or compostable. Furthermore, some biodegradable films may also be unduly sensitive to water, either limiting the use of the film or requiring some type of surface treatment to the film, often rendering the film nonbiodegradable. Others have inadequate heat resistance for wide spread use. Thus, there is a need for a film which is completely biodegradable.

The use of lactic acid and lactide to manufacture a biodegradable polymer is known in the medical industry. As disclosed by Nieuwenhuis et al. (U.S. Patent No. 5,053,485), such polymers have been used for making biodegradable sutures, clamps, bone plates and biologically active controlled release devices. Processes developed for the manufacture of polymers to be utilized in the medical industry have incorporated techniques which respond to the need for high purity and biocompatibility in the final product. These processes were designed to produce small volumes of high dollar-value products, with less emphasis on manufacturing cost and yield.

In order to meet projected needs for biodegradable packaging materials, others have endeavored to optimize lactide polymer processing systems. Gruber et al. (U.S. Pat. No. 5,142,023) disclose a continuous process for the manufacture of lactide polymers with controlled optical purity from lactic acid having physical properties suitable for replacing present petrochemical-based polymers.

Generally, manufacturers of polymers utilizing processes such as those disclosed by Gruber et al. will convert raw material monomers into polymer beads, resins or other pelletized or powdered products. The polymer in this form may then be sold to end users who convert, i.e., extrude, blow-mold, cast films, blow films, thermoform, injection-mold or fiber-spin the polymer at elevated temperatures to form useful articles. The above processes are collectively referred to as melt-processing. Polymers produced by processes such as those disclosed by Gruber et al., which are to be sold commercially as beads, resins, powders or other non-finished solid forms are generally referred to collectively as polymer resins.

Prior to the present invention and its related patent applications, it is believed that there has been no disclosure of a combination of composition control and melt stability requirements which will lead to the production of commercially viable, amorphous and semi-crystalline lactide polymer films.

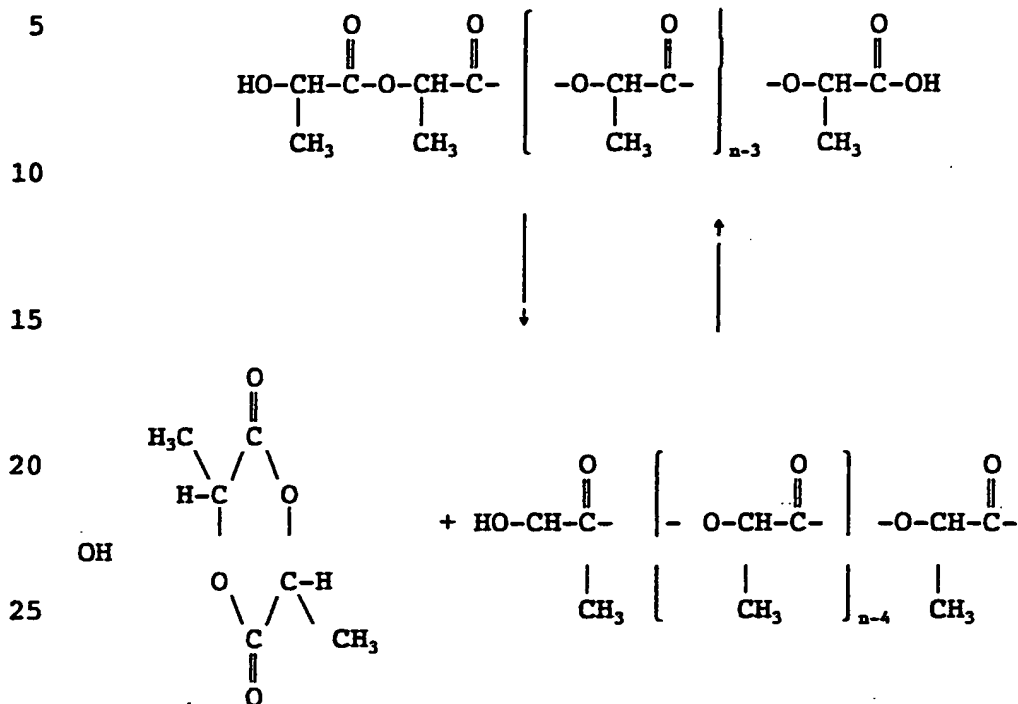
It is generally known that lactide polymers or poly(lactide) are unstable. The concept of instability has both negative and positive aspects. A positive aspect is the biodegradation or other forms of degradation which occur when lactide polymers or articles manufactured from lactide polymers are discarded or composted after completing their useful life. A negative aspect of such instability is the degradation of lactide polymers during processing at

elevated temperatures as, for example, during melt-processing by end-user purchasers of polymer resins. Thus, the same properties that make lactide polymers desirable as replacements for non-degradable

- 5 petrochemical polymers also create undesirable effects during processing which must be overcome.

Lactide polymer degradation at elevated temperatures has been the subject of several studies, including: I. C. McNeill and H. A. Leiper, Polymer Degradation and
10 Stability, vol. 11, pp. 267-285 (1985); I. C. McNeill and H. A. Leiper, Polymer Degradation and Stability, vol. 11, pp. 309-326 (1985); M. C. Gupta and V. G. Deshmukh, Colloid & Polymer Science, vol. 260, pp. 308-311 (1982); M. C. Gupta and V. G. Deshmukh, Colloid &
15 Polymer Science, vol. 260, pp. 514-517 (1982); Ingo Luderwald, Dev. Polymer Degradation, vol. 2, pp. 77-98 (1979); Domenico Garozzo, Mario Giuffrida, and Giorgio Montaudou, Macromolecules, vol. 19, pp. 1643-1649 (1986); and, K. Jamshidi, S. H. Hyon and Y. Ikada, Polymer, vol.
20 29, pp. 2229-2234 (1988).

It is known that lactide polymers exhibit an equilibrium relationship with lactide as represented by the reaction below:



No consensus has been reached as to what the primary degradation pathways are at elevated processing temperatures. One of the proposed reaction pathways includes the reaction of a hydroxyl end group in a "back-biting" reaction to form lactide. This equilibrium reaction is illustrated above. Other proposed reaction pathways include: reaction of the hydroxyl end group in a "back-biting" reaction to form cyclic oligomers, chain scission through hydrolysis of the ester bonds, an intramolecular beta-elimination reaction producing a new acid end group and an unsaturated carbon-carbon bond, and radical chain decomposition reactions. Regardless of the mechanism or mechanisms involved, the fact that substantial degradation occurs at elevated temperatures, such as those used by melt-processors, creates an obstacle to

use of lactide polymers as a replacement for petrochemical-based polymers. It is apparent that degradation of the polymer during melt-processing must be reduced to a commercially acceptable rate while the polymer maintains the qualities of biodegradation or compostability which make it so desirable. It is believed this problem has not been addressed prior to the developments disclosed herein and in related disclosures.

As indicated above, poly(lactide)s have been produced in the past, but primarily for use in medical devices. These polymers exhibit biodegradability, but also a more stringent requirement of being bioresorbable or biocompatible. As disclosed by M. Vert, Die Ingwandte Makromolekulare Chemie, vol. 166-167, pp. 155-168 (1989), "[t]he use of additives is precluded because they can leach out easily in body fluids and then be recognized as toxic, or, at least, they can be the source of fast aging with loss of the properties which motivated their use. Therefore, it is much more suitable to achieve property adjustment through chemical or physical structure factors, even if aging is still a problem." Thus, work aimed at the bioresorbable or biocompatible market focused on poly(lactide) and blends which did not include any additives.

Other disclosures in the medical area include Nieuwenhuis (European Patent No. 0 314 245), Nieuwenhuis (U.S. Patent No. 5,053,485), Eitenmuller (U.S. Patent No. 5,108,399), Shinoda (U.S. Patent No. 5,041,529), Fouty (Canadian Patent No. 808,731), Fouty (Canadian Patent No. 923,245), Schneider (Canadian Patent No. 863,673), and Nakamura et al., Bio. Materials and Clinical Applications, Vol. 7, p. 759 (1987). As disclosed in these references, in the high value, low volume medical specialty market, poly(lactide) or lactide polymers and copolymers can be given the required physical properties by generating lactide of

very high purity by means of such methods as solvent extraction or recrystallization followed by polymerization. The polymer generated from this high purity lactide is a very high molecular weight product which will retain its physical properties even if substantial degradation occurs, and the molecular weight drops significantly during processing. Also, the polymer may be precipitated from a solvent in order to remove residual monomer and catalysts. Each of these treatments adds stability to the polymer but clearly at a high cost, which would not be feasible for lactide polymer compositions that are to be used to replace inexpensive petrochemical-based polymers in the manufacture of films.

Furthermore, it is well-known that an increase in molecular weight generally results in an increase in a polymer's viscosity. A viscosity which is too high can prevent melt-processing of the polymer due to physical/mechanical limitations of the melt-processing equipment. Melt-processing of higher molecular weight polymers generally requires the use of increased temperatures to sufficiently reduce viscosity so that processing can proceed. However, there is an upper limit to temperatures used during processing. Increased temperatures increase degradation of the lactide polymer, as the previously-cited studies disclose.

Jamshidi et al., Polymer, Vol. 29, pp. 2229-2234 (1988) disclose that the glass transition temperature of a lactide polymer, T_g , plateaus at about 57°C for poly(lactide) having a number average molecular weight of greater than 10,000. It is also disclosed that the melting point, T_m , of poly (L-lactide) levels off at about 184°C for semi-crystalline lactide polymers having a number average molecular weight of about 70,000 or higher. This indicates that at a relatively low molecular weight, at least some physical properties of lactide polymers plateau and remain constant.

Sinclair et al. (U.S. Patent No. 5,180,765) disclose the use of residual monomer, lactic acid or lactic acid oligomers to plasticize poly(lactide) polymers, with plasticizer levels of 2-60 percent. Loomis (U.S. Patent 5 No. 5,076,983) discloses a process for manufacturing a self-supporting film in which the oligomers of hydroxy acids are used as plasticizing agents. Loomis and Sinclair et al. disclose that the use of a plasticizer such as lactide or oligomers of lactic acid is 10 beneficial to produce more flexible materials which are considered to be preferable. Sinclair et al., however, disclose that residual monomer can deposit out on rollers during processing. Loomis also recognizes that excessive levels of lactide or oligomers of lactic acid 15 can cause unevenness in films and may separate and foul processing equipment. Thus, plasticizing as recommended, negatively impacts melt-processability in certain applications.

Accordingly, a need exists for a lactide polymer 20 which is melt-stable under the elevated temperatures common to melt-processing resins in the manufacture of film. The needed melt-stable polymer composition must also exhibit sufficient compostability or degradability after its useful life as a film. Further, the melt-stable polymer must be processable in existing melt- 25 processing equipment, by exhibiting sufficiently low viscosities at melt-processing temperatures while polymer degradation and lactide formation remains below a point of substantial degradation and does not cause 30 excessive fouling of processing equipment. Furthermore, the lactide polymer must retain its molecular weight, viscosity and other physical properties within commercially-acceptable levels through film manufacturing processes. It will be further appreciated 35 that a need also exists for methods of manufacturing such films. The present invention addresses these needs as well as other problems associated with existing

lactide polymer compositions and manufacturing processes. The present invention also offers further advantages over the prior art, and solves other problems associated therewith. ^{opt}

5

SUMMARY OF THE INVENTION

According to the present invention, amorphous and semi-crystalline biodegradable films are provided. The films comprise a melt-stable, lactide polymer
10 composition comprising: poly(lactide) polymer chains having a number average molecular weight of at least 10,000; the polymer being a reaction product of polymerizing a lactide mixture comprising meso-lactide. For an amorphous film the lactide mixture comprises more
15 than about 1 percent by weight meso-lactide, and the remaining lactide can be L-lactide, D-lactide or mixtures thereof. The polymer, because it is amorphous, exhibits a net melting endotherm of less than about 10 joules per gram of poly(lactide). For a semi-
20 crystalline film the lactide mixture comprises less than about 15 percent by weight meso-lactide, and the remaining lactide can be L-lactide, D-lactide or mixtures thereof provided that the overall lactide mixture comprises at least about 85% of either the L or
25 D lactide isomer. Both the amorphous and semi-crystalline polymers have residual lactide in a concentration of less than about 2 percent by weight and optionally water in a concentration of less than about 2,000 parts per million. Processes for the manufacture
30 of the films are also provided. For the purposes of the present invention, the films may be manufactured from any number of methods and are not to be limited by any particular method.

Optionally, stabilizing agents in the form of anti-
35 oxidants and water scavengers may be added. Further, plasticizers, nucleating agents, anti-static agents, slip aids and anti-blocking agents may be added. The

resultant films are biodegradable and may be disposed of in an environmentally sound fashion.

The present invention recognizes conditions and compositions which affect the crystallization of poly(lactide) films. In particular, it is shown that in certain cases an amorphous film may become semi-crystalline. Compositions which assure an amorphous, non-crystallizing film, are pointed out.

The present invention describes a method of increasing the degree of crystallinity in a film or sheet by drawing the film in a machine and/or transverse direction orientation at temperatures near the T_g . The T_g can be lowered to near room temperature through the use of plasticizers.

Strain hardening is a phenomenon which, if present, can be used to obtain high quality, uniform, semi-crystalline films. A description of strain hardening in stretching of films of poly(ethylene 2,6, naphthalene dicarboxilate) is given by Cakmak et al. [M. Cakmak, Y.D. Wang, and M. Simhambhatla, Polymer Engineering and Science, June 1990, Vol. 30, p 721-733]. Strain hardening can be identified by an increase in the force required to continue elongation of a film. The essential feature of this phenomenon is the appearance of necks (thin areas) during the stretching operation. Larger amounts of stretching occur locally in the necked region, causing it to elongate more than the surrounding areas. The elongation presumably causes the further crystallization of the previous "weak" neck. A neck elsewhere in the film then forms, elongates, crystallizes, hardens and increases its resistance to further elongation. This continues until all areas of the film have once again reached a uniform thickness. As shown by Cakmak et al., the process results in very smooth, high quality films. We have observed necking in poly(lactide) as it has been subjected to stress induced crystallization, and believe that strain hardening may

be occurring. The films, which have stretched, often feel smoother to the touch, although no surface profiling tests have been done. Poly(lactide) is a polymeric material which offers unique advantages as a
5 film not only in the biodegradable sense, but in the manufacturing process as well.

The films of the present invention may be used in articles such as diapers, packaging film, shrink wrap film, agricultural mulch film, bags and tape. These
10 films are superior in diaper constructions as compared to other films such as polypropylene or polyethylene. The typical construction of a diaper comprises an outer, water impervious back sheet, a middle absorbent layer and an inner layer. The outer back sheet, comprising
15 the exterior of the diaper, is desirable from an environmental point of view if it is biodegradable. The films of the present invention satisfy this environmental concern by being biodegradable and compostable.

20 Furthermore, a poly(lactide) film, unlike other biodegradable polymers, is believed not to support microbial growth during storage and typical film use. Starch or other biodegradable polymers, when exposed to warm, damp environments, will promote the growth of
25 unhealthy microbes. This is undesirable in personal hygiene products. Thus, the present invention has yet another advantage over prior biodegradable polymers.

Another advantage of the present invention is the high surface energy of poly(lactide) films.
30 Poly(lactide) is a material with a relatively high surface energy, when compared to other films. As the surface energy of an extruded film increases, the driving force to remain intact and to minimize surface area increases, therefore the tendency to form a smooth,
35 coherent, high gloss film increases. A high surface energy film also has the advantage of having a surface

on which it is easier to print. This is an important feature in packaging applications and diapers.

A film of the present invention exhibits a higher surface energy than untreated polyolefin films. In order to produce a satisfactory printing surface, polyolefin films must first be modified. Modification not only increases the costs associated with production of the films, but the modification treatment will diffuse into the film and produce an unsatisfactory printing surface.

The surface energy of substantially pure poly(lactide) films of the present invention is about 44 dynes/cm. This leads to a surface with satisfactory printing characteristics without surface modification. Slip aids or other additives may reduce the surface energy down to about 35 dynes/cm. Additionally, inks which are typically more difficult to apply onto paper coatings, like water based inks, may be applied directly to poly(lactide).

Poly(lactide) is a relatively low viscosity polymer which¹ allows the extrusion of the film to be done at lower temperatures than traditional films. This results in a cost savings to the converter because the extrusion equipment will not require as much power when run at lower temperatures.

Heat sealability is also a property of films which is desirable. Poly(lactide) can be heat sealed at temperatures lower than 70°C, at line pressures lower than 40 psi, and at times less than 2 sec.

30

Polymer Blends

To improve certain properties of poly(lactide), it may be advantageous to blend a second polymer with poly(lactide). The polymer chosen for blending with poly(lactide) will be one which has the properties necessary for the particular need and is compatible with poly(lactide) to the extent that the particular

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properties of poly(lactide) are improved.

Incompatibility often results in a polymer blend which has inferior properties, such as very low tensile strength, rheological properties, degradability, and barrier properties to oxygen, moisture or carbon dioxide. Polymers which may be useful for improving the film properties of poly(lactide) include aliphatic polyesters or polyamides made by both ring opening and condensation polymerization, esterified cellulose resins, derivitized starch, polyvinylacetate and any of its partially hydrolyzed products including polyvinylalcohol, polyethers including poly(ethylene oxide), polycarbonates, polyurethanes including those based on aliphatic isocyanates, polyanhydrides, natural rubber and its derivatives including epoxidized natural rubber, block copolymers of styrene and isoprene or butadiene and the hydrogenated version of those polymers, polyacrylates and methacrylates, polyolefins, and polystyrene.

Examples of particular interest include polymers which are also degradable including poly(caprolactone), poly(hydroxybutyrate hydroxyvalerate), cellulose acetate, cellulose acetate butyrate, cellulose acetate propionate, and poly(vinyl alcohol).

These polymers may be blended with poly(lactide) in percentages of about 1 to 95% by weight to make films of improved properties as shown in Example 1.

The above described features and advantages along with various other advantages and features of novelty are pointed out with particularity in the claims of the present application. However, for a better understanding of the invention, its advantages, and objects attained by its use, reference should be made to the drawings which form a further part of the present application and to the accompanying descriptive matter in which there is illustrated and described preferred embodiments of the present invention.

BRIEF DESCRIPTION OF THE DRAWINGS

In the drawings, in which like reference numerals indicate corresponding parts or elements of preferred
5 embodiments of the present invention throughout the several views;

Fig. 1 is a schematic representation of a preferred process for the manufacture of a melt-stable lactide polymer composition; and

10 Fig. 2 is a graph showing the equilibrium relationship between lactide and poly(lactide) at various temperatures.

Fig. 3 is a graph showing the melting endotherm for annealed samples of poly(lactide).

15 Fig. 4 is a phase diagram for meso-lactide, L-lactide and D-lactide.

DETAILED DESCRIPTION OF THE PREFERRED EMBODIMENT(S)

The lactide polymer compositions used in films
20 disclosed herein focus on meeting the requirements of the melt-processor of a lactide polymer resin such as that produced from a process disclosed by Gruber et al. in U.S. Patent 5,142,023. However, the present invention is directed to a poly(lactide) film and is not
25 limited to the lactide polymer composition or process of Gruber et al. Any lactide polymer composition, which comes within the scope of this invention, may be used as a film. As disclosed herein, the problems of degradation, fouling, and lactide formation during melt-
30 processing of lactide polymers are addressed through suggested ranges of molecular weights and compositional limits on impurities, such as residual monomer, water and catalyst along with the use of stabilizing agents and catalyst-deactivating agents.

35 In general, according to the present invention, melt-stable lactide polymer films and processes for manufacturing melt-stable lactide polymer film from a

melt-stable lactide polymer are disclosed. The use of the term "film" includes films and sheets. Lactide polymers are useful due to their recyclable and biodegradable nature. Furthermore, lactide polymers are
5 compostable as illustrated in Example 16 below.

Applicants believe the hydrolysis of the ester may be the key to or the first step in degradation of a lactide polymer composition. The mechanism of degradation is not key to the films of the present invention, however
10 it must be recognized that such degradation makes lactide polymers desirable as replacements for presently-utilized non-degradable petrochemical-based polymers used for films.

Applicants have found that the instability of
15 lactide polymers, which leads to the beneficial degradation discussed above, also creates processing problems. These processing problems include generation of lactide monomer at elevated temperatures and loss in molecular weight. Applicants believe that the loss in
20 molecular weight is due to chain scission degradation of the ester bonds and other depolymerization reactions which are not completely understood. No consensus has been reached as to what the primary degradation pathways at elevated processing temperatures are. As previously
25 disclosed, these pathways may include equilibrium-driven depolymerization of lactide polymers to form lactide and chain scission through hydrolysis of the ester bonds along with other pathways. For purposes of the present invention, the exact mechanism of degradation at
30 elevated temperatures is not critical.

It is to be understood, however, that degradation of lactide polymers is both beneficial and detrimental. Benefits derive from degradability when articles
manufactured from such polymers are discarded. The same
35 or similar types of degradation are detrimental if they occur during processing or prior to the end of the article's useful life.

Lactic acid has two optical isomers, L-lactic acid, also known as (S)-lactic acid, and D-lactic acid, also known as (R)-lactic acid. Three forms of lactide can be derived from the two forms of lactic acid. They are

5 L,L-lactide, also known as L-lactide and which comprises two (S)-lactic acid residuals; D,D-lactide, also known as D-lactide and which comprises two (R)-lactic acid residuals; and meso-lactide, which comprises one each of (R)- and (S)-lactic acid residuals. A 50/50 said

10 mixture of D-lactide and L-lactide with a melting point of about 126°C is sometimes called D,L-lactide. At temperatures higher than the melting point, it is essentially a liquid mixture of D-lactide and L-lactide.

The similarities and differences between poly(lactic

15 acid) and various poly(lactide)s can best be examined by looking at the distribution of (R) and (S)-lactic acid residuals in the polymers. An L-lactide or D-lactide will introduce a pair of (S) or (R) residuals into the chain, respectively. Meso-lactide introduces an (R,S)

20 or (S,R) dyad. The characteristics of the final polymer will depend for various applications, on the sequencing of the (R) and (S) residuals.

Crystallinity requires relatively long sequences of a particular residual, either long sequences of (R) or

25 of (S). The length of the interrupting sequences may be important for establishing other features of the polymer, such as the rate at which it crystallizes or the melting point of the crystalline phase, or melt processability. The table below shows the expected

30 statistical distribution of the major and minor sequence lengths assuming random polymerization and neglecting transesterification. The table shows data for mixtures containing predominately the (S) configuration, the same results would be obtained for mixtures containing

35 predominately the (R) configuration.

<u>Monomer mix</u>	<u>Optical Composition</u>	<u>Probability major sequence having length at least n=</u>			<u>Probability minor sequence having length at least n=</u>		
		6	10	20	1	2	3
<u>Lactic Acid</u>							
95 L / 5 D	95 S / 5 R	0.77	0.63	0.38	1.0	0.05	0.002
90 L / 10 D	90 S / 10 R	0.59	0.39	0.14	1.0	0.10	0.01
<u>Lactide</u>							
95 L / 5 D	95 S / 5 R	0.90	0.81	0.63	1.0	1.0	0.05
90 L / 10 D	90 S / 10 R	0.81	0.66	0.39	1.0	1.0	0.10
85 L / 15 D	85 S / 15 R	0.72	0.52	0.23	1.0	1.0	0.15
80 L / 20 D	80 S / 20 R	0.64	0.41	0.13	1.0	1.0	0.20
<u>Meso mixtures</u>							
95 L / 5 meso	97 S / 3 R	0.90	0.81	0.63	1.0	0.01	0.00
90 L / 10 meso	95 S / 5 R	0.81	0.66	0.39	1.0	0.02	0.00
85 L / 15 meso	92 S / 8 R	0.72	0.52	0.23	1.0	0.04	0.00
80 L / 20 meso	90 S / 10 R	0.64	0.41	0.13	1.0	0.05	0.00

The table above shows, that for the L-lactide system, D-lactide or meso-lactide result in similar major sequence lengths at similar levels. The major sequence length is believed to dominate whether or not crystallization can occur. Fischer et al. [Fischer, E.W., Sterzel, H.J., and Wegner, G., Kolloid-Z. u. Z Polymere 251, p980-990 (1973)] studied the system of L-lactide and D-lactide and reported that crystallization did not occur if the minor component was more than 15% of the polymerization mixture. Our results, documented in Example 24, show that polymers made of L-lactide and meso-lactide will not crystallize when the polymerization mixture contains more than about 15% of the meso-lactide. These results are consistent with the table above, and suggest that a lactide or lactic acid polymer is crystallizable provided that there is at least a 0.5 probability that sequences of the major conformation comprise at least 10 lactic acid residuals.

The table above also shows that polymers of predominately L-lactide with either D-lactide or meso-lactide as minor components have dramatically different sequences of the minor component. For polymers made of L-lactide and meso-lactide there is no chance of having three or more (R)-lactic acid residuals in a row, and a very low probability of having two in a row. For polymers made of L-lactide with low concentrations of D-lactide, the (R)-lactic acid residuals always appear in at least a sequence of length two, with a significant fraction appearing as sequences of length four.

Polymers made either from L- and D-lactic acid (by direct condensation, for example) or from L-lactide with small amounts of meso-lactide have a somewhat similar structure when compared at similar levels of (S) and (R) residuals, as shown in the table above.

Melt-Processing

It is believed that a manufacturer of lactide polymers from a lactide monomer will produce a lactide

polymer resin which is in the form of beads or pellets. The melt-processor will convert the resin to a film by elevating the temperature of the resin above at least its glass transition temperature but normally higher and extruding the resin into a film. It is to be understood that the conditions of elevated temperature used in melt-processing cause degradation of lactide polymers during processing. Degradation under melt-processing conditions is shown experimentally in Example 8 based on equilibrium, Example 11 based on catalyst concentration, Example 12 based on catalyst activity, Example 14 based on use of stabilizers and Example 15 based on moisture content. As can be seen in these examples, it is understood that several factors appear to affect the rate of degradation during melt-processing. Applicants have addressed these factors in a combination of compositional requirements and the addition of stabilizing or catalyst-deactivating agents to result in a polymer of lactide which is melt-stable.

In addition, melt-processing frequently produces some proportion of trimmed or rejected material. Environmental concerns and economical efficiencies dictate that this material be reused, typically by regrinding and adding back the material into the polymer feed. This introduces additional thermal stress on the polymer and increases the need for a melt-stable polymer composition.

Melt Stability

The lactide polymers of the present invention are melt-stable. By "melt-stable" it is meant generally that the lactide polymer, when subjected to melt-processing techniques, adequately maintains its physical properties and does not generate by-products in sufficient quantity to foul or coat processing equipment. The melt-stable lactide polymer exhibits reduced degradation and/or reduced lactide formation relative to known lactide polymers. It is to be

understood that degradation will occur during melt-processing. The compositional requirements and use of stabilizing agents as disclosed herein reduce the degree of such degradation to a point where physical properties are not significantly affected by melt-processing, and fouling by impurities or degradation by-products such as lactide does not occur. Furthermore, the melt-stable polymer should be melt-processable in melt-processing equipment such as that available commercially. Further, the polymer will preferably retain adequate molecular weight and viscosity. The polymer should preferably have sufficiently low viscosity at the temperature of melt-processing so that the extrusion equipment may create an acceptable film. The temperature at which this viscosity is sufficiently low will preferably also be below a temperature at which substantial degradation occurs.

Polymer Composition

The melt-stable lactide polymer film of the present invention comprises a poly(lactide) polymer chains having a number average molecular weight of at least about 10,000 and preferably from about 10,000 to about 300,000. In a preferred composition for a film, the number average molecular weight ranges from about 20,000 to about 275,000. In the most preferred composition, the number average molecular weight ranges from about 40,000 to about 250,000.

The molecular weight of a polymer sample can be determined through the use of gel permeation chromatography (GPC). In the present case, the GPC analysis was conducted with an Ultrastyrigel[®] column from Waters Chromatography. The mobile phase was chloroform. A refractive index detector with molecular weight calibration using polystyrene standards was used. The GPC temperature was 35°C. Molecular weights were determined by integrating from the highest molecular weight fraction down to 4,000 amu. The region below

4,000 amu is excluded from the calculations of molecular weight in order to improve reproducibility of the number average molecular weight. This material may be separately reported as "oligomers" and residual lactide, as in Example 12.

In the present invention, a film is considered to be semi-crystalline if it exhibits a net melting endotherm of greater than about 10 J/gm of poly(lactide) when analyzed by a differential scanning calorimeter DSC. To determine whether a film is semi-crystalline it can be tested in a differential scanning calorimeter (DSC), such as marketed by Mettler. An accurately weighed sample of the film, weighing between 5 mg and 15 mg, is placed in the test ampule. A suitable temperature program is to start at -20°C and scan at 20°C/min to 200°C. Typical features which may be observed include a glass transition at a temperature designated T_g , a relaxation endotherm peak immediately following T_g , a crystallization exotherm peak (generally in the range of 70-140°C), and a melting endotherm peak (generally in the range of 100-200°C). The net melting endotherm is the energy of the melting endotherm less the energy of the crystallization exotherm if present.

As detailed in Example 10, it appears that the physical properties such as modulus, tensile strength, percentage elongation at break, impact strength, flexural modulus, and flexural strength remain statistically constant when the lactide polymer samples are above a threshold molecular weight. As detailed in Example 23, there is a practical upper limit on molecular weight based on increased viscosity with increased molecular weight. In order to melt-process a high molecular weight lactide polymer, the melt-processing temperature must be increased to reduce the viscosity of the polymer. As pointed out in the Examples, the exact upper limit on molecular weight must be determined for each melt-processing application in

that required viscosities vary and residence time within the melt-processing equipment will also vary. Thus, the degree of degradation in each type of processing system will also vary. Based on the disclosure of Example 23,
5 it is believed that one could determine the suitable molecular weight upper limit for meeting the viscosity and degradation requirements in any application.

Lactide polymers can be in either an essentially amorphous form or in a semi-crystalline form. For
10 various applications it will be desirable to have the polymer in semi-crystalline form. Semi-crystalline films have superior heat resistance. The tendency of films to adhere together at temperatures experienced during manufacture, use, shipping or storage when on a
15 roll or part of a product is reduced for semi-crystalline films. Semi-crystalline films also have decreased permeation to gases, such as oxygen, and moisture. This is an advantage in packaging applications, especially food packaging. Lactide
20 polymer films with increased crystallinity generally degrade more slowly than amorphous films under conditions of high humidity and heat which results in extended shelf life of the films.

The desired range of compositions for semi-
25 crystalline poly(lactide) is less than about 15 percent by weight meso-lactide and the remaining percent by weight being either L-lactide or D-lactide, wherein at least 85 percent comprises either the L or D-lactide isomer. A more preferred composition contains less than
30 about 12 percent by weight meso-lactide and a most preferred composition has less than about 9 percent by weight meso-lactide with the remainder being substantially all L-lactide and/or D-lactide.

Addition of even small amounts of meso-lactide to
35 the polymerization mixture results in a polymer which is slower to crystallize than polymerization mixtures having lesser amounts of meso-lactide, as detailed in

Example 24. Beyond about 15 percent meso content the polymer remains essentially amorphous following the annealing procedure of Example 26.

There are four main methods to increase the rate of crystallization. One is to increase chain mobility at low temperatures, by adding, for example, a plasticizing agent. Dioctyl adipate is an example of a plasticizer which helps crystallization rates in poly(lactide), as detailed in Example 27. A second method to increase the rate of crystallization is to add a nucleating agent, as detailed in Example 28. A third method to induce crystallinity is to orient the polymer molecules. Orientation can be accomplished by drawing during film casting, blowing films, stretching a sheet after it is cast or blown (in multiple directions, if desired), or by the flow of polymer through a small opening in a die. When the process of orientation results in crystallization it is known as stress induced crystallization. This phenomena is illustrated for poly(lactide) in Examples 32 and 33. A fourth method of inducing crystallization is heat-setting, which involves holding a constrained oriented film or fiber at temperatures above T_g . It is demonstrated in Example 29. Heat setting involves exposing the film to elevated temperatures, as shown in *Plastics Extrusion Technology*, F. Hensen (ed), Hanser Publishers, New York, 1988, pp 308, 324. It is preferred to heat set the film under tension to reduce shrinkage during the setting process.

It has been found that poly(lactide) having a meso-content of less than about 12% may be drawn just above its T_g in a machine direction orientation (MDO) or transverse direction orientation (TDO) process to increase the degree of crystallinity. In instances where the T_g of the composition is below room temperature, such as compositions containing at least 15% plasticizer, the sheet may be drawn at room temperature to increase levels of crystallinity from

less than 5 J/gm to greater than 15 J/gm. Example 32 demonstrates the increase in crystallinity of a plasticized poly(lactide) sheet upon drawing. The properties of the crystallized and plasticized film are superior with regard to flexible film over non crystallized film. Crystallizing a plasticized film increases the blocking temperature of the film as shown in Example 33. The tensile strength and barrier properties will also increase upon crystallization.

Crystallizing lactide polymer films may be performed by drawing the film in either the machine direction or transverse direction or in both directions using draw ratios of 1.1 or greater. The temperature of the draw rolls are generally set at temperatures at or slightly above the T_g of the film.

The degree of crystallinity in lactide polymer films containing at least 15% plasticizer may also be increased by storing the film at room temperature until levels of crystallinity greater than 10 J/g is reached. Storing the film under elevated temperatures may increase the rate of crystallization, especially in lactide polymer films containing less than 15% plasticizer.

Crystallization of the lactide polymer may also be done during the manufacture of resin pellets. The crystalline portions of the polymer are melted during film manufacture, therefore recrystallization during film manufacture is still required from semi-crystalline films. However, crystalline resin pellets may be dried at higher temperatures, therefore faster than amorphous resin pellets due to the increased resistance of semi-crystalline resin pellets to adhere together at elevated temperatures. Crystallization of the resin pellets may be done by drawing the strand of polymer to a draw ratio of at least 1.1 as it exits the extruder and prior to being pelletized. Crystallinity may also be increased in lactide polymers containing at least 15% plasticizer

by storing the pellets at room temperature for a period of time necessary to increase crystallinity above 10 Joules per gram.

Crystalline poly L-lactide exhibits an endotherm of
5 roughly 92 joules per gram at its melting temperature of 170°-190°C, as shown by S. Gogolewski and A.J. Pennings, J. Applied Polymer Science, Vol. 28, pp 1045-1061 (1983). The melting point changes with composition. The degree of crystallinity is roughly proportional to
10 the endotherm on melting. For purposes of the present invention, semi-crystalline poly(lactide) exhibits a net melting endotherm above about 10 joules per gram of poly(lactide).

In the present invention, one of the films
15 comprising a melt-stable lactide polymer in a preferred embodiment is substantially amorphous. To determine whether a film is amorphous, it can be tested in a differential scanning calorimeter (DSC), such as marketed by Mettler. The process of determining whether
20 a film is amorphous is substantially the same as disclosed previously for determining whether a film is semi-crystalline. In the present invention, a film is considered to be amorphous if it exhibits a net melting endotherm of less than about 10 J/gm of poly(lactide).
25 The net melting endotherm is the energy of the melting endotherm less the energy of the crystallization exotherm, if present.

As detailed in Examples 36, 37 and 38, an amorphous polymer potentially may crystallize during a hydrolysis
30 test, thereby losing its quality of being amorphous. This test is believed to represent a critical step in the composting of poly(lactide). As detailed in Example 16, amorphous polymers exhibit superior degradability when subjected to a compost test.

35 In one preferred embodiment, the film is non-crystallizing, in addition to being amorphous. A non-crystallizing film is one which does not develop

crystallinity when subjected to a proscribed annealing procedure, such as detailed in Example 26. As detailed in Example 26, a non-crystallizing polymer can be made from a lactide composition containing at least about 12 percent by weight of meso-lactide. Figure 4 illustrates a preferred embodiment, showing a non-crystallizing polymer needing to have at least 15 percent by weight meso-lactide when the other component is either pure L-lactide or pure D-lactide. However, Example 26 illustrates the meso-lactide concentration can be as low as above 12 percent by weight and still be non-crystallized. If the remaining lactide is not substantially optically pure, then less meso-lactide would be required to prevent crystallization. Meso-lactide concentrations can be as low as 1% and still result in an amorphous polymer. In addition to using the composition of the poly(lactide) to ensure non-crystallizing of the polymer, copolymers, blends or other structural modifications to the polymer could be used to prevent crystallization in a preferred embodiment.

The residual monomer concentration in the melt-stable lactide polymer composition is less than about 2.0 percent by weight. In a preferred composition, the lactide concentration is less than about 1.0 percent by weight and a most preferred composition has less than about 0.5 percent by weight of lactide. Contrary to disclosures in the art, Applicants have found that the monomer cannot be used as a plasticizing agent in the resin of the present invention due to significant fouling of the extrusion equipment. As detailed in Example 17, it is believed the low levels of monomer concentration do not plasticize the final polymer.

The water concentration within the melt-stable lactide polymer composition is as low as reasonably possible and preferably less than about 2,000 parts-per-million. Preferably this concentration is less than 500

parts-per-million and most preferably less than about 100 parts-per-million. As detailed in Example 15, the polymer melt-stability is significantly affected by moisture content. Thus, a melt-stable polymer of the present invention must have the water removed prior to melt-processing. Applicants recognize that water concentration may be reduced prior to processing the polymerized lactide to a resin. Thus, moisture control could be accomplished by packaging such resins in a manner which prevents moisture from contacting the already-dry resin. Alternatively, the moisture content may be reduced at the melt-processor's facility just prior to the melt-processing step in a dryer. Example 15 details the benefit of drying just prior to melt-processing and also details the problems encountered due to water uptake in a polymer resin if not stored in a manner in which moisture exposure is prevented or if not dried prior to melt-processing. As detailed in these examples, Applicants have found that the presence of water causes excessive loss of molecular weight which may affect the physical properties of the melt-processed polymer.

In a preferred composition of the present invention, a stabilizing agent is included in the polymer formulation to reduce degradation of the polymer during production, devolatilization, drying and melt processing by the end user. The stabilizing agents recognized as useful in the present films may include antioxidants and/or water scavengers. Preferred antioxidants are phosphite-containing compounds, hindered phenolic compounds or other phenolic compounds. The antioxidants include such compounds as trialkyl phosphites, mixed alkyl/aryl phosphites, alkylated aryl phosphites, sterically hindered aryl phosphites, aliphatic spirocyclic phosphites, sterically hindered phenyl spirocyclics, sterically hindered bisphosphonites, hydroxyphenyl propionates, hydroxy benzyls, alkylidene

bisphenols, alkyl phenols, aromatic amines, thioethers, hindered amines, hydroquinones and mixtures thereof. As detailed in Example 14, many commercially-available stabilizing agents have been tested and fall within the scope of the present melt-stable lactide polymer film. Biodegradable antioxidants are particularly preferred.

The water scavengers which may be utilized in preferred embodiments of the melt-stable lactide polymer film include: carbodiimides, anhydrides, acyl chlorides, isocyanates, alkoxy silanes, and desiccant materials such as clay, alumina, silica gel, zeolites, calcium chloride, calcium carbonate, sodium sulfate, bicarbonates or any other compound which ties up water. Preferably the water scavenger is degradable or compostable. Example 20 details the benefits of utilizing a water scavenger.

In a preferred composition of the present invention, a plasticizer is included in the polymer formulation to improve the film quality of the lactide polymer. More particularly, plasticizers reduce the melt viscosity at a given temperature of poly(lactide). This reduction assists in processing and extruding the polymer at lower temperatures, may improve film flexibility, reduces cracking tendencies of the finished film, and improves impact and tear resistance of the film.

A plasticizer is useful in concentration levels of about 1 to 40 percent based on the weight of polymer. Preferably, a plasticizer is added at a concentration level of about 5 to 25 percent. Most preferably, a plasticizer is added at a concentration level of about 8 to 25 percent.

Selection of a plasticizing agent requires screening of many potential compounds and consideration of several criteria. For use in a biodegradable film the preferred plasticizer is to be biodegradable, non-toxic and compatible with the resin and relatively nonvolatile.

Plasticizers in the general classes of alkyl or aliphatic esters, ether, and multi-functional esters and/or ethers are preferred. These include alkyl phosphate esters, dialkylether diesters, tricarboxylic
5 esters, epoxidized oils and esters, polyesters, polyglycol diesters, alkyl alkylether diesters, aliphatic diesters, alkylether monoesters, citrate esters, dicarboxylic esters, vegetable oils and their derivatives, and esters of glycerine. Most preferred
10 plasticizers are tricarboxylic esters, citrate esters, esters of glycerine and dicarboxylic esters. Citroflex A4[®] from Morflex is particularly useful. These esters are anticipated to be biodegradable. Plasticizers containing aromatic functionality or halogens are not
15 preferred because of their possible negative impact on the environment.

For example, appropriate non-toxic character is exhibited by triethyl citrate, acetyltriethyl citrate, tri-n-butyl citrate, acetyltri-n-butyl citrate,
20 acetyltri-n-hexyl citrate, n-butyltri-n-hexyl citrate and dioctyl adipate. Appropriate compatibility is exhibited by acetyltri-n-butyl citrate and dioctyl adipate. Other compatible plasticizers include any plasticizers or combination of plasticizers which can be
25 blended with poly(lactide) and are either miscible with poly(lactide) or which form a mechanically stable blend. Corn oil and mineral oil were found to be incompatible when used alone with poly(lactide) because of phase separation (not mechanically stable) and migration of
30 the plasticizer.

Volatility is determined by the vapor pressure of the plasticizer. An appropriate plasticizer must be sufficiently non-volatile such that the plasticizer stays substantially in the resin formulation throughout
35 the process needed to produce the film. Excessive volatility can lead to fouling of process equipment, which is observed when producing films by melt

processing poly(lactide) with a high lactide content. Preferred plasticizers should have a vapor pressure of less than about 10 mm Hg at 170°C, more preferred plasticizers should have a vapor pressure of less than
5 10 mm Hg at 200°C. Lactide, which is not a preferred plasticizer, has a vapor pressure of about 40 mm Hg at 170°C. Example 6 highlights useful plasticizers for the present invention.

Internal plasticizers, which are bonded to the
10 poly(lactide) may also be useful. Epoxides provide one method of introducing an internal plasticizer.

In a preferred composition, nucleating agents may be incorporated during polymerization. Nucleating agents may include selected plasticizers, finely divided
15 minerals, organic compounds, salts of organic acids and imides and finely divided crystalline polymers with a melting point above the processing temperature of poly(lactide). Examples of useful nucleating agents include talc, sodium salt of saccharin, calcium
20 silicate, sodium benzoate, calcium titanate, boron nitride, copper phthalocyanine, isotactic polypropylene, crystalline poly(lactide) and polybutylene terephthalate.

In a preferred composition, fillers may be used to
25 prevent blocking or sticking of layers or rolls of the film during storage and transport. Inorganic fillers include clays and minerals, either surface modified or not. Examples include talc, diatomaceous earth, silica, mica, kaolin, titanium dioxide, perlite, and
30 wollastonite. Preferred inorganic fillers are environmentally stable and non-toxic.

Organic fillers include a variety of forest and agricultural products, either with or without modification. Examples include cellulose, wheat,
35 starch, modified starch, chitin, chitosan, keratin, cellulosic materials derived from agricultural products, gluten, nut shell flour, wood flour, corn cob flour, and

guar gum. Preferred organic fillers are derived from renewable sources and are biodegradable. Fillers may be used either alone or as mixtures of two or more fillers. Examples 4 and 5 highlight useful anti-blocking fillers for the present invention. Surface treatments such as corona and flame treatments may also be used to reduce blocking.

Pigments or color agents may also be added as necessary. Examples include titanium dioxide, clays, calcium carbonate, talc, mica, silica, silicates, iron oxides and hydroxides, carbon black and magnesium oxide.

In the manufacture of the melt-stable lactide polymer compositions of the present invention, the reaction to polymerize lactide is catalyzed. Many catalysts have been cited in literature for use in the ring-opening polymerization of lactones. These include but are not limited to: SnCl_2 , SnBr_2 , SnCl_4 , SnBr_4 , aluminum alkoxides, tin alkoxides, zinc alkoxides, SnO , PbO , Sn (2-ethyl hexanoates), Sb (2-ethyl hexanoates), Bi (2-ethyl hexanoates), Na (2-ethyl hexanoates) (sometimes called octoates), Ca stearates, Mg stearates, Zn stearates, and tetraphenyltin. Applicants have also tested several catalysts for polymerization of lactide at 180°C which include: tin(II) bis(2-ethyl hexanoate) (commercially available from Atochem, as Fascat 2003, and Air Products as DABCO T-9), dibutyltin diacetate (Fascat 4200[®], Atochem), butyltin tris(2-ethyl hexanoate) (Fascat 9102[®], Atochem), hydrated monobutyltin oxide (Fascat 9100[®], Atochem), antimony triacetate (S-21, Atochem), and antimony tris(ethylene glycoxide) (S-24, Atochem). Of these catalysts, tin(II) bis(2-ethyl hexanoate), butyltin tris(2-ethyl hexanoate) and dibutyltin diacetate appear to be most effective.

Applicants have found the use of catalysts to polymerize lactide significantly affects the stability of the resin product. It appears the catalyst as incorporated into the polymer also is effective at

catalyzing the reverse depolymerization reaction.

Example 11 details the effect of residual catalyst on degradation. To minimize this negative effect, in a preferred composition, the residual catalyst level in the resin is present in a molar ratio of initial monomer-to-catalyst greater than about 3,000:1, preferably greater than about 5,000:1 and most preferably greater than about 10,000:1. Applicants believe a ratio of about 20,000:1 may be used, but polymerization will be slow. Optimization of catalyst levels and the benefits associated therewith are detailed in Example 21. Applicants have found that when the catalyst level is controlled within these parameters, catalytic activity is sufficient to polymerize the lactide while sufficiently low to enable melt-processing without adverse effect when coupled with low residual monomer level and low water concentration as described above in polymers of molecular weight of at least about 10,000 and preferably between about 10,000 to about 300,000. It is believed in most applications the addition of a stabilizing agent may be unnecessary if the catalyst level is optimized.

Applicants have also found that catalyst concentration may be reduced subsequent to polymerization by precipitation from a solvent. Example 22 demonstrates potential catalyst removal by precipitation from a solvent. This produces a resin with reduced catalyst concentration. In an alternative embodiment, the catalyst means for catalyzing the polymerization of lactide to form the poly(lactide) polymer chains which was incorporated into the melt-stable lactide polymer composition during polymerization is deactivated by including in the melt-stable lactide polymer composition a catalyst deactivating agent in amounts sufficient to reduce catalytic depolymerization of the poly(lactide) polymer chains. Example 12 details the benefits of utilizing a catalyst deactivating agent.

Such catalyst-deactivating agents include hindered, alkyl, aryl and phenolic hydrazides, amides of aliphatic and aromatic mono- and dicarboxylic acids, cyclic amides, hydrazones and bishydrazones of aliphatic and aromatic aldehydes, hydrazides of aliphatic and aromatic mono- and dicarboxylic acids, bis-acylated hydrazine derivatives, and heterocyclic compounds. A preferred metal deactivator is Irganox® MD1024 from Ciba-Geigy. Biodegradable metal deactivators are particularly preferred.

In an alternative embodiment, the catalyst concentration is reduced to near zero by utilizing a solid-supported catalyst to polymerize lactide. The feasibility of utilizing such a catalyst is detailed in Example 9. It is believed catalysts which may be utilized include supported metal catalysts, solid acid catalysts, acid clays, alumina silicates, alumina, silica and mixtures thereof.

In a preferred composition, the catalyst usage and/or deactivation is controlled to reduce depolymerization of the poly(lactide) polymer during melt-processing to less than about 2 percent by weight generation of lactide from a devolatilized sample in the first hour at 180°C and atmospheric pressure. More preferably, the amount of lactide generated is less than about 1 percent by weight in the first hour and most preferably less than about 0.5 percent by weight in the first hour.

A preferred melt-stable lactide polymer composition is the reaction product of polymerization of lactide at a temperature greater than about 160°C. Applicants have found that polymerization at higher temperatures result in a characteristically different polymer which is believed to have improved melt stability due to increased transesterification during polymerization. The benefits of higher temperature polymerization are detailed in Example 13.

Melt-Stable Lactide Polymer Process

The process for the manufacture of a melt-stable lactide polymer comprises the steps of first providing a lactide mixture wherein the mixture contains less than
5 15 percent by weight meso-lactide with the remainder being L-lactide and/or D-lactide. Such purified lactide stream may be such as that produced in the process disclosed by Gruber et al., although the source of lactide is not critical to the present invention.

10 The lactide mixture is polymerized to form a lactide polymer or poly(lactide) with some residual unreacted monomer in the presence of a catalyst means for catalyzing the polymerization of lactide to form poly(lactide). Catalysts suitable for such
15 polymerization have been listed previously. The concentration of catalysts utilized may be optimized as detailed in the following examples and discussed previously.

In a preferred embodiment, a stabilizing agent,
20 which may be an antioxidant and/or a water scavenger is added to the lactide polymer. It is recognized that such stabilizing agents may be added simultaneously with or prior to the polymerization of the lactide to form the lactide polymer. The stabilizing agent may also be
25 added subsequent to polymerization.

As previously disclosed, the catalyst usage is adjusted and/or deactivation agent is added in a sufficient amount to reduce depolymerization of poly(lactide) during melt-processing to less than 2
30 percent by weight generation of lactide from a devolatilized sample in the first hour at 180°C and atmospheric pressure. More preferably, the stabilizing agent controls lactide generation to less than 1 percent by weight and most preferably less than 0.5 percent by
35 weight in the first hour at 180°C and atmospheric pressure. Alternatively, the control of catalyst concentration to optimize the balance between necessary

catalytic activity to produce poly(lactide) versus the detrimental effects of catalytic depolymerization or degradation of the lactide polymer may be utilized to obviate the need for adding a stabilizing agent.

5 The lactide polymer is then devolatilized to remove unreacted monomer which may also be a by-product of decomposition reactions or the equilibrium-driven depolymerization of poly(lactide). Any residual water which may be present in the polymer would also be
10 removed during devolatilization, although it is recognized that a separate drying step may be utilized to reduce the water concentration to no greater than about 2,000 parts-per-million. The devolatilization of the lactide polymer may take place in any known
15 devolatilization process. The key to selection of a process is operation at an elevated temperature and usually under conditions of vacuum to allow separation of the volatile components from the polymer. Such processes include a stirred tank devolatilization or a
20 melt-extrusion process which includes a devolatilization chamber and the like. An inert gas sweep is useful for improved devolatilization.

In a preferred process for manufacture of a melt-stable lactide polymer composition, the process also
25 includes the step of adding a molecular weight control agent to the lactide prior to catalyzing the polymerization of the lactide. For example, molecular weight control agents include active hydrogen-bearing compounds, such as lactic acid, esters of lactic acid,
30 alcohols, amines, glycols, diols and triols which function as chain-initiating agents. Such molecular weight control agents are added in sufficient quantity to control the number average molecular weight of the poly(lactide) to between about 10,000 and about 300,000.

35 Next referring to Figure 1 which illustrates a preferred process for producing a melt-stable lactide polymer composition. A mixture of lactides enters a

mixing vessel (3) through a pipeline (1). A catalyst for polymerizing lactide is also added through a pipeline (13). Within mixing vessel (3) a stabilizing agent may be added through a pipeline (2). A water
5 scavenger may also be added through the pipeline (2). The stabilized lactide mixture is fed through a pipeline (4) to a polymerization process (5). The polymerized lactide or lactide polymer leaves the polymerization process through a pipeline (6). The stream is fed to a
10 second mixing vessel (8) within which a stabilizing agent and/or catalyst deactivating agent may be added through a pipeline (7). The stabilized lactide polymer composition is then fed to a devolatilization process (10) through a pipeline (9). Volatile components leave
15 the devolatilization process through a pipeline (11) and the devolatilized lactide polymer composition leaves the devolatilization process (10) in a pipeline (12). The devolatilized lactide composition is fed to a resin-finishing process (14). Within the resin-finishing
20 process the polymer is solidified and processed to form a pelletized or granular resin or bead. Applicants recognize the polymer may be solidified and processed to form resin or bead first, followed by devolatilization. The resin is then fed to a drying process (16) by
25 conveyance means (15). Within the drying process (16) moisture is removed as a vapor through pipeline (17). The dried lactide polymer resin leaves the drying process (16) by a conveyance means (18) and is fed to a melt-processing apparatus (19). Within the melt-
30 processing apparatus (19) the resin is converted to a useful article as disclosed above. The useful article leaves the melt-processing apparatus (19) through a conveyance means (20).

The following examples further detail advantages of
35 the system disclosed herein:

Example 1Blended Film

Polycaprolactone commercially available as TONE 787 from Union Carbide was mixed with poly(lactide) having a number average molecular weight of 157,900, a residual lactide concentration of 0.19%, a meso-lactide concentration of about 10% and a water concentration of less than about 500 ppm on a Leistritz twin screw extruder at 12.8%, 25.6%, 38.4% by weight to poly(lactide). These blends were injection molded into standard ASTM test bars using a Cincinnati Milacron Vista Sentry VST-55 molding press, and physical properties were measured and tensiles tested on the bars. The above blends were also extruded into cast film on a Killion extruder with a sheet die, die gap of 0.035" and with 0.25% by weight Celite Super Floss diatomaceous earth used as an anti-block agent.

The following table illustrates critical data:

20

Table 1

Compounding: Twin Screw
Conditions

Cast Film: Killion

conditions

25 zone 1: 150°C

zones 2-8: 180°C

zones 9 & 10: 170°C

melt temp: 194°C

die pressure: 44 psi

30 amps: 17.8-20.0

screw rpm: 300

Pressure in

devolatization zone

200 mm hg

zone 1: 286°F

zone 2: 320°F

zone 3: 320°F

zone 4: 315°F

adaptor: 310°F

die temp: 310°F

pressure: 1340 psi

screw rpm: 20.2

35

The following table illustrates the results. It is noted that blends of poly(caprolactone) and poly(lactide) are more flexible than unblended poly(lactide) as shown by the increased elongation. This is significant for flexible films as well as other products where decreased brittleness is required.

Table 2

<u>Film:</u>	<u>Mn:</u>	<u>Mw:</u>	<u>PDI:</u>	<u>% Lactide</u>	<u>Break Str. (kpsi)</u>	<u>Modulus (kpsi)</u>
Example 1 DVD 97	66662	157866	2.37	0.19	7.9	317
Example 2 12.8% tone	68564	170697	2.49	0.33	8.6	350
Example 3 25.6% tone	95355	240973	2.53	0.08	N/A	N/A
Example 4 38.6% tone & 10% citro.A-4s	87013	221583	2.55	0.36	N/A	N/A

<u>Injection Molded Test Bars:</u>	<u>Break Str. (kpsi)</u>	<u>Modulus (kpsi)</u>	<u>Elongation at break (%)</u>
Example 6 DVD-97	7.2	493	5.6 +-1.9
Example 7 12.7% tone	6.6	403	227 +-52
Example 8 25.6% tone	3.2	340	250 +-133
Example 9 38.6% tone	2.7	332	149 +-58

Example 2Example Showing Orientation of Poly(lactide) for Film

To demonstrate the ease and benefit of making oriented poly(lactide) film, an experiment was done to make a sheet using poly(lactide) and orientating that sheet to different degrees. The polymer selected had a 280,000 to 400,000 weight average molecular weight, a 100,000 to 150,000 number average molecular weight, lactide concentration lower than 1%, meso level of about 10 to 20%, and a moisture level lower than about 500 ppm.

The sheet was cast using a 2" diameter single barrier flight screw Davis-Standard extruder with a 24:1 L/D. The die was 24" wide with an adjustable gap. The casting roll was 30" wide and 24" in diameter and equipped with temperature controls. After casting, the sheet was oriented in both the machine (MD) and transverse (TD) directions.

The MD orienter consisted of six 30" wide rolls and five temperature control zones: roll 1 and 2 were for preheat, roll 3 for slow draw, roll 4 for fast draw, roll 5 for heat setting, and roll 6 for cooling the sheet. Draw ratios up to 20 were used for the MD orientation.

The TD was a standard tenter frame with a 3 zone oven. The feed width range for the tenter was 8-48" and the output width range was 16-62". Following the orientation section was a slitter/winder.

Typical conditions used in extruding the sheet are shown below:

Extruder zone 1	350 to 400°F	(177-204°C)
zone 2	360 to 400°F	(182-204°C)
zone 3	340 to 350°F	(171-188°C)
Melt pip/adapters	330 to 350°F	(166-177°C)
Die zones	330°F	
Head pressure	1640 to 2280 psi	
Screw speed	45 to 75 rpm	

Melt strength at the die was very good at 330°F (166°C). The die was positioned about 0.5 inch off the cast roll, which was run between 112 and 126°F. An air
5 knife was used to pin the melt against the cast roll to reduce neck-in.

Conditions found useful for machine direction orientation of poly(lactide) were temperatures of 65°C for the preheat rolls, 65°C and 72°C for the slow draw
10 roll at low draw ratios and high draw ratios respectively, 65°C for the fast draw roll, 45°C for the heat roll, and 20°C for the cooling roll.

The gap between the slow and fast roll was set to the minimum possible. Orientation took place only
15 slightly above the T_g to give a high degree of molecular alignment. Rolls were collected after MD orientation and some were used for TD orientation.

Conditions for transverse direction orientation were 63°C for the preheat zone, 70°C for the stretching zone,
20 and 67°C for the annealing zone. Ambient air was circulated at the oven exit to cool the oriented sheet before winding.

The poly(lactide) oriented very well, being easily curved over the rolls and requiring lower process
25 temperatures than standard plastic. The products made were:

	<u>MD</u>	<u>TD</u>	<u>Approx.</u>
30 <u>Thickness</u>			
Sample 1	0	0	5 mil
Sample 2	3.2	0	5 mil
Sample 3	3.5	4.3	1 mil
Sample 4	3.5	2.0	2 mil
35 Sample 5	1.5	2.9	9 mil
Sample 6	1.5	2.0	13 mil

Samples 1, 2 and 3 were tested for tensile and elongation according to ASTM D882 and tear resistance
40 according to ASDTM D689. The results are shown below:

Table 3

		<u>Tensile Strength at break (lb/in/mil)</u>	<u>Elongation at break (%)</u>	<u>Tear Strength (g-cm)</u>
5				
	Sample 1	6.9	4.3	40
10		7.4	4.8	40
	Sample 2	15.8	85.1	16
		6.9	3.2	40
15	Sample 3	23.9	89.3	32
		10.5	160	128

As the data in the table shows, the tensile strength and flexibility of the poly(lactide) film can be greatly increased by orientation.

Example 3Non-Blocking Characteristics of Plasticized, Crystallized Poly(lactide)

Poly(lactide) having a weight average molecular weight of 165,000, a meso level of about 6%, a lactide level of less than about 0.2%, and a moisture level of less than about 500 ppm was blended with 25% by weight acetyl tributylcitrate (Citroflex A4 from Morflex) in a Werner & Pfleiderer twin screw extruder.

Two films of this composition were prepared in a Carver press equipped with heated platens at a temperature of 300-350°F and a dwell time of about 60 seconds. These films were annealed in an oven over four days to induce crystallization of the poly(lactide) samples. The films were tested for resistance to blocking at an elevated temperature. This test was performed by placing two films in contact with each other in an oven held at 60°C with a 95 gram weight with a surface area of about 2.5 in² on top of the two films. After more than four hours, the films were removed from the oven and any adhesion between the films was noted. No adhesion occurred during the test. Prior to the

annealing and crystallization procedure, films adhered to one another at room temperature.

This result shows that poly(lactide) films containing high levels of plasticizer will have adequate blocking resistance once the poly(lactide) is crystallized.

Example 4

Anti-Block Aids

10 The use of anti-block aids can increase the resistance of two poly(lactide) films to stick together at elevated temperatures. This was demonstrated using poly(lactide) having a weight molecular weight of 165,000, a residual lactide level of about 0.1%, a meso
15 level of about 10%, and a moisture level of about 60 ppm. The anti-block aid was diatomaceous earth having a median particle size of 3.5 microns (Celite Super Floss from Celite) which was dried to a moisture level of about 400 ppm. The diatomaceous earth and poly(lactide)
20 were blended in a twin screw extruder at different levels of anti-block aid and pelletized. The pellets were converted into film using the single screw extruder as in Example 2. The films were tested for resistance to adhering to one another by placing two films together
25 with a 92 gram weight on top in an oven set at 60°C for 2 hours. A failure was when the films could not be separated after being removed from the oven. The results are shown in the table below:

Table 4

<u>Sample I.D.</u>	<u>% Celite Super Floss</u>	<u>Blocking Test 1</u>	<u>Blocking Test 2</u>
2113-90-0	0.0	Fused	Fused
2113-90-1	4.1	None	None
2113-90-2	7.9	Some tiny pts. stuck	None
2113-90-3	1.8	Some pts. stuck	None
2113-90-4	0.9	Some pts. stuck	None
2113-90-5	0.45	Some pts. stuck	None

Example 5Anti-Blocking Agents

Two injection molded disks, 2.5 inch diameter, were placed together with a 94 gram weight on top and held at 50°C for 24 hours. The disks had the following agents compounded therein. The disks were then cooled to room temperature and pulled apart by hand and ranked for blocking characteristics (considerable, slight and none). The following are the results:

10

Table 5

<u>AGENTS</u>		
15	Poly(lactide) control	considerable
	22% wheat gluten	none
	10% wheat gluten	slight
20	22% pecan shell	none
	15% pecan shell	slight
	23% wollastonite	slight
25	28% Ultratalc 609	none
	23% Ultratalc 609	none
	28% Microtuff F talc	slight
	22% Microtuff F talc	slight
30	14% Microtuff F talc	slight
	2% Microtuff F talc	considerable

Example 6

35

Plasticizer Agents

Dried pellets of devolatilized poly(lactide) were processed in a twin screw extruder to allow compounding of various plasticizing agents. The strands leaving the extruder were cooled in a water trough and chopped into pellets. Samples of the pellets were heated at 20°C/minute to 200°C in a DSC apparatus, held at 200°C for 2 minutes and rapidly cooled to quench the samples. The quenched samples were then reheated in the DSC apparatus increasing at 20°C/minute to determine the

45

glass transition temperature. These samples were compared to a polymer with no plasticizer. The effect of the plasticizer on the glass transition temperature is shown in the table below. Glass transition temperatures are taken at the mid-point of the transition.

Table 6

	<u>SAMPLE</u>	<u>T_g(C)</u>	<u>Change in T_g/wt. percent additive</u>
10	Control	54.8	---
	8% Dioctyl adipate	35.0	2.5
	Control+40% silica	54.5	---
15	Control+40% silica+		
	5% dioctyl adipate	36.0	3.7
	Control	54.6	---
	6% Citroflex A-4*	42.6	2.0
20	12% Citroflex A-4	31.4	1.9
	Control	59.3	---
	1.6% Citroflex A-4	56.3	1.9
	2.9% Citroflex A-4	53.1	2.1
25	Control	58.4	---
	2.1% Citroflex A-4	56.1	1.1
	3.4% Citroflex A-4	50.5	2.3
30	Control	61.6	---
	18.6% Citroflex A-2	54.7	0.4
	13.1% Citroflex B-6	52.4	0.7
	12.6% Citroflex A-6	53.8	0.6

*Citroflex is a registered trademark of Morflex, Inc., Greensboro, NC. A-4 is the designation of a purified acetyltri-n-butyl citrate. A-2 is the designation of acetyltriethyl citrate, A-6 is the designation of acetyltri-n-hexyl citrate, and B-6 is the designation of n-butyryltri-n-hexyl citrate.

These results show the effectiveness of these plasticizers in reducing the glass transition temperature of poly(lactide).

The procedure above was tried using corn oil as a plasticizer. Visual observation showed the corn oil to be not compatible, forming a film on the surface. Corn

oil and mineral oil were both not effective as a primary plasticizer with poly(lactide). They may still be useful as a secondary plasticizer, in combination with a compatible primary plasticizer.

5

Example 7

Heat Sealability of a Poly(lactide) Film:

Poly(lactide) with a weight average molecular weight of 160,000, a meso-lactide content of 8 to 12%, and a
10 moisture level of about 200 ppm, was extruded into a film using the same method as Example 2.

Two pieces of the poly(lactide) film were heat sealed in an Accuseal Heat Sealer Model 30 having heated platens of a dimension of 1/4"x25". The temperature and
15 dwell time were varied to find the minimum required to form a bond between the films. Pressure was 40 psi. Bond strength was estimated by pulling the two films apart 10 seconds after sealing and judging pass or fail. Either the films tore before the bond failed (pass) or
20 the films separated at the bondline (fail). Results are shown below:

Table 7

		<u>Temperature (C°)</u>									
25	Time (sec)	30	40	50	60	65	70	75	80	85	90
	1	F	F								
	2			F	F	F	P	P			
30	3				F				P	P	
	4					P					P5
	P										
	6										P

35 Sealing conditions of 70°C at a line pressure of 40 psi for a 2 second dwell time were found to be sufficient for forming a good bond between the two poly(lactide) films.

Example 8Lactide and Poly(lactide) Equilibrium Concentrations

Experiments were conducted to determine the equilibrium concentration of lactide and poly(lactide) at different temperatures. In these experiments a sample of lactide was polymerized in the presence of a catalyst (tin (II) bis(2-ethyl hexanoate)) and held at a fixed temperature for 18 hours or greater. Beyond this time the residual monomer concentration is believed essentially constant. The content of residual monomer was determined by GPC analysis. GPC analysis was conducted with an Ultrastyrigel® column from Waters Chromatography. The mobile phase was chloroform. A refractive index detector with molecular weight calibration using polystyrene standards was used. The GPC temperature was 35°C. Data analysis was completed using the software package Baseline, model 810, version 3.31.

The results of tests conducted on several samples at various temperatures are summarized in the graph of Fig. 2 as indicated by X's on such graph. Also plotted on the graph of Fig. 2 are data points cited in A. Duda and S. Penczek, Macromolecules, vol. 23, pp. 1636-1639 (1990) as indicated by circles on the graph. As can be seen from the graph of Fig. 2, the equilibrium concentration, and thus the driving force behind the depolymerization of poly(lactide) to form lactide, increases dramatically with increased temperature. Thus, melt-processing at elevated temperatures results in degradation of the lactide polymer to form lactide on the basis of equilibrium alone. For example, lactide concentrations below about 2 percent cannot be directly obtained at temperatures of 140°C or above due to the identified equilibrium relationship between lactide and poly(lactide).

Example 9Lactide Polymerization in the Presence
of a Solid Supported Catalyst5 Tin (II) Oxide

24 grams of L-lactide (melting point about 97°C) and
6 grams of D,L-lactide (for the purposes of this
invention, D,L-lactide has a melting point of about
126°C) were combined in a round bottom flask with 0.033
10 grams of Tin (II) oxide, as a fine powder. This
corresponds to the catalyst level of 852:1, molar ratio
lactide to tin. The flask was then purged with dry
nitrogen 5 times. This was lowered into an oil bath at
160°C with magnetic stirring. Polymerization time was 8
15 hours.

Amberlyst 36

24 grams of L-lactide and 6 grams of D,L-lactide
were combined in a round bottom flask with 1.06 grams of
20 Amberlyst 36 resin beads. The flask was purged 5 times
with dry nitrogen. The flask was lowered into an oil
bath at 140°C with magnetic stirring. Polymerization
time was 8 hours. The resin had a stated proton content
of 1 meq/gram dry weight resin. The resin was prepared
25 by rinsing 2 times with 10 volumes dry methanol, then
dried for several hours under high vacuum for several
hours at 40°C.

The polymerization results are shown below:

30

TABLE 8

	<u>Sample</u>	<u>Mn</u>	<u>Mw</u>	<u>PDI</u>	<u>% Conversion</u>
35	Tin (II) Oxide	77,228	103,161	1.34	54.0
	Amberlyst	1,112	1,498	1.34	73.5

40

Example 10Molecular Weight Relationship to Physical
Properties of Lactide Polymers

Poly(lactide) samples with various molecular weights and optical compositions were prepared by polymerizing blends of L-lactide and meso-lactide at 180°C under nitrogen in a 1-gallon sealed reactor. Tin(II) bis(2-ethyl hexanoate) catalyst was added at a monomer-to-catalyst ratio of 10,000:1. After about 1 hour the molten polymer was drained from the reactor using nitrogen pressure. The sample was poured into a pan and placed in a vacuum oven at about 160°C for about 4 hours to bring the reaction to near equilibrium levels.

Portions of the samples were then dried under vacuum and processed in an injection molding apparatus (New Britain 75 from New Britain Machine Co.) to produce standard test bars for physical property testing. The results of physical property testing are shown in the following Table 9. The physical property tests were made according to ASTM methods D 638, D 256, and D 790. The reported results are the averages of several tests.

Samples of the test bars after injection molding were analyzed by GPC for molecular weight. Other portions of the test bars were reground and tested in a capillary viscometer to determine the melt-viscosity. These results are also included in Table 9.

Statistical analysis of the data revealed no correlations which were statistically significant between either optical composition or molecular weight and the mechanical properties of modulus, tensile strength, percentage elongation at break, notched Izod impact strength, flexural modulus, or flexural strength. The independence of these properties on molecular weight indicates that all of these samples were above a "threshold" molecular weight required to achieve the intrinsic properties of the polymer in a preferred composition.

The viscosity data show significant correlations with molecular weight. This dependence documents the practical limitation and necessity of controlling polymer molecular weight below an upper
5 limit at which it is impractical to melt-process the polymer. At high molecular weight, high viscosity prevents processing by standard melt-processing equipment. Increases in temperature to reduce viscosity dramatically increase polymer degradation and lactide
10 formation which is also unacceptable.

TABLE 9

Sample I.D.	Meso Lactide In Blend, Wt%	Molecular Weight After Injection	Viscosity at 173°C (Pa-S)		
			Final IV (dl/g)	Shear Rate 100 S ⁻¹	Shear Rate 1000 S ⁻¹
6	40	41000	0.86	5.5	2.9
5	10	54000	0.88	10.4	7.2
4	20	59000	0.91	10.4	7.2
8	10	64000	1.02	15.7	10.0
9	40	68000	0.97	12.6	8.1
7	20	71000	1.16	36.0	12.9
10	20	83000	1.19	35.8	15.8

Mechanical Properties of Injection Molded Samples

Sample I.D.	Modulus MPa	Tensile Strength (Yld) PSI	% Elongation at Break	Flexural IZOD Impact ft · lb./in	Flexural Modulus MPa	Strength PSI
6	0.55	6600	3.3	0.39	0.53	11300
5	0.56	7800	3.5	0.46	0.54	12500
4	0.56	7600	3.9	0.32	0.53	12500
8	0.55	7700	3.4	0.47	0.53	12400
9	0.59	6700	3.1	0.42	0.52	10600
7	0.56	7400	3.3	0.45	0.51	12400
10	0.55	6700	3.0	0.47	0.52	9900

Example 11Effect of Residual Catalyst on Polymer Degradation

Polymer samples were prepared at four levels of catalyst, corresponding to monomer to catalyst molar ratios of 5,000:1, 10,000:1, 20,000:1, and 40,000:1. The catalyst utilized was tin (II) bis(2-ethyl hexanoate). These samples were then subjected to heating in a TGA apparatus (TA Instruments, Inc., model 951 thermogravimetric analyzer with a DuPont 9900 computer support system) with a nitrogen purge. Isothermal conditions of 200°C for 20 minutes were used. The samples were then analyzed by GPC with a viscosity-based detector and a universal calibration curve to determine the extent of breakdown in molecular weight. The GPC apparatus for this test was a Viscotek Model 200 GPC and a Phenomenex column. The TGA analysis typically resulted in about a 5 percent loss in weight and molecular weight drops of 0 to 70 percent.

The number average molecular weights were converted to a milliequivalent per kilogram basis (1,000,000/Mn) in order to calculate a rate of chain scission events. The results below represent averages of 2-4 replicates on each of the four samples.

TABLE 10

	<u>Catalyst level</u> <u>(monomer/catalyst)</u>	<u>Scission Rate</u> <u>(meq/kg*min)</u>
30	5,000	1.33
	10,000	0.62
	20,000	0.44
	40,000	0.12

The rate of chain scission was directly proportional to the residual catalyst level, demonstrating the detrimental effect of catalyst activity on melt-stability under conditions similar to melt-processing. This instability, however, is distinguished from the instability due to the equilibrium relationship between

lactide and poly(lactide) detailed in Example 8, in that loss of molecular weight due to catalytic depolymerization by chain scission is evident.

5

Example 12Catalyst Deactivation Experiment

Two runs were made in a laboratory Parr reactor. Lactide feed was 80 percent L-lactide and 20 percent D,L-lactide. Molecular weight was controlled by adding
10 a small quantity of lactic acid, the target molecular weight was 80,000 Mn.

Lactide was charged to the reactor as a dry mix, the reactor was purged 5 times with nitrogen, and heated up to 180°C. At this point catalyst (5000:1 initial
15 monomer to catalyst molar ratio, Fascat® 2003) was charged through a port in the top of the reactor. The reaction was allowed to proceed for 70 minutes at 180°C, with mechanical agitation. Conversion at this point was
20 93-94 percent, close to the equilibrium value at 180°C of 96 percent poly(lactide) from Figure 2. This point is considered t-zero, designating the completion of the polymerization reaction and the beginning of the mixing time.

In the control experiment, a sample was taken and
25 the mixture was held at temperature with continued agitation. Samples were taken periodically through a port in the reactor bottom. After 4 hours the reactor was drained.

In the example experiment, a sample was taken and
30 0.25 weight percent of a metal deactivator (Irganox® MD 1024®) was added through the catalyst addition port. The mixture was held at temperature with continued agitation and samples were withdrawn periodically. The reactor was drained after 4 hours.

35 GPC analysis (utilizing the method of Example 8) for these samples was divided into three parts: polymer with molecular weight over 4,000 (for which the Mn and

Mw numbers are reported), the percent oligomers (comprising the region with molecular weight greater than lactide but less than 4,000, as distinguished from oligomers as defined by Loomis to include only oligomers up to a molecular weight of 450), and percent lactide (residual monomer). The structure of the oligomers was not certain, but it is believed they were primarily cyclic structures. It is also believed that the metal deactivator, if unreacted, will elute with the oligomer fraction. Quantification of the oligomer fraction is difficult, because the GPC trace is near the baseline in this region.

The analysis of the polymer samples as withdrawn from the reactor at various time intervals for the control and experimental compositions are shown below in Table 11.

TABLE 11

20	Control	Mn	Mw	% Polymer	% Oligomer	% Monomer
	t-zero	67,100	119,500	94	0	6.0
	0.5 hr	62,500	119,000	95	0.7	3.9
	1.0 hr	61,500	116,100	96	0	3.6
25	1.5 hr	56,000	111,600	95	1.5	3.3
	2.0 hr	57,600	110,900	96	0.9	3.1
	4.0 hr	51,400	105,400	94	3.3	3.1
30	Test	Mn	Mw	% Polymer	% Oligomer	% Monomer
	t-zero	63,200	110,700	93	3.5	3.8
	0.5 hr	52,100	108,600	92	4.6	2.9
	1.0 hr	52,700	109,200	92	4.9	2.8
35	1.5 hr	53,400	107,200	93	4.0	3.1
	2.0 hr	59,700	111,100	94	0.6	5.8
	4.0 hr	51,200	107,300	91	6.1	3.3

40 The samples were then ground and placed in a 120°C oven under vacuum (pressure 0.1 inch Hg) for 14 hours. Sample analyses after this treatment are shown below in Table 12.

TABLE 12

	Control	Mn	Mw	% Polymer	% Oligomer	% Monomer
5	t-zero	45,500	88,500	98	2.2	0.0
	0.5 hr	45,000	88,700	98	2.0	0.0
	1.0 hr	43,900	87,200	98	2.0	0.0
	1.5 hr	42,600	84,000	98	2.2	0.0
	2.0 hr	42,000	85,200	97	3.2	0.0
10	4.0 hr	41,900	82,800	98	2.0	0.0
	Test	Mn	Mw	% Polymer	% Oligomer	% Monomer
	t-zero	39,300	76,700	96	4.0	0.0
15	0.5 hr	43,900	85,100	98	2.4	0.0
	1.0 hr	55,300	98,600	96	3.8	0.0
	1.5 hr	48,400	96,200	95	4.5	0.0
	2.0 hr	48,900	101,900	95	5.0	0.0
20	4.0	50,600	101,900	94	5.6	0.0

- In all cases the polymer was completely devolatilized (0.0 percent residual lactide monomer). The data also clearly show that the metal deactivator
- 25 reduced the degradation of polymer during the devolatilization step (as indicated by the greater loss in Mn for the control samples from Table 11 to Table 12 versus the Test samples). One hour of mixing appears to be long enough to develop most of the benefit.
- 30 The samples were stored at room temperature under nitrogen for about 1 week and reanalyzed, as shown below in Table 13.

TABLE 13

	Control	Mn	Mw	Z Polymer	Z Oligomer	Z Monomer
5	t-zero	33,500	71,000	100	0.1	0.0
	0.5 hr	43,400	95,800	99	1.0	0.0
	1.0 hr	44,900	96,300	100	0.1	0.0
	1.5 hr	45,900	95,000	100	0.0	0.0
	2.0 hr	45,900	94,100	100	0.2	0.0
10	4.0 hr	43,100	90,100	99	1.3	0.0

	Test	Mn	Mw	Z Polymer	Z Oligomer	Z Monomer
15	t-zero	44,600	84,900	100	0.0	0.0
	0.5 hr	45,300	90,600	99	1.2	0.0
	1.0 hr	47,800	100,000	98	2.4	0.0
	1.5 hr	46,600	98,900	96	3.5	0.0
20	4.0	57,700	110,200	96	4.0	0.3

Equilibrium lactide levels are estimated to be less than 0.2 weight percent at room temperature. Consistent with that, essentially no lactide was observed in any of the samples (detection limit about 0.1 weight percent). The oligomer content in the non-stabilized samples declined and some increase in molecular weight was noted, perhaps due to reincorporation of the (cyclic) oligomers into the polymer. The oligomer depletion reaction was inhibited in the stabilized polymers, with the extent of inhibition dependent on the length of time that the additive was mixed.

The samples were then reheated to 180°C in sealed vials and held for one hour as a simulation of melt-processing. Analysis of the samples after the heat treatment is given below in Table 14.

TABLE 14

	Control	Mn	Mw	% Polymer	% Oligomer	% Monomer
	t-zero	23,900	60,000	88	8.4	4.0
5	0.5 hr	23,900	59,600	90	7.7	2.7
	1.0 hr	23,700	58,800	88	9.3	2.7
	1.5 hr	24,700	58,000	86	10.0	3.8
	2.0 hr	26,100	56,400	90	6.8	2.7
10	4.0 hr	24,800	58,700	92	6.6	1.9

	Test	Mn	Mw	% Polymer	% Oligomer	% Monomer
	t-zero	33,900	64,300	95	2.2	3.1
15	0.5 hr	17,900	34,600	94	4.8	1.7
	1.0 hr	21,200	42,900	94	4.6	1.8
	1.5 hr	29,200	56,900	98	0.5	1.8
	2.0 hr	missing				
20	4.0 hr	35,700	71,400	95	3.7	1.7

The data for molecular weight show that if the metal deactivator is not mixed into the system long enough then it can have a detrimental impact on stability in the melt. The samples for which the mixing was at least 1.5 hours show no detrimental effect, and the 4 hour sample appears to be somewhat more stable than any of the others based on molecular weight alone. More importantly, the metal deactivator samples show significantly less lactide reformation than the control samples. This effect is gained even in the samples which were mixed for only 0.5 hour. The metals deactivated samples averaged only 1.8 percent lactide after one hour at 180°C, compared to an average of 3.0 percent lactide for the controls. The equilibrium level at 180°C is about 3.6 percent from Figure 2. Thus, the use of metal deactivators can reduce the troublesome reformation of lactide during melt-processing of the finished polymer.

40

Example 13Effect of Increased Polymerization Temperature
on Polymer Characteristics

L-lactide (Boeringer Ingelheim, S-grade) was used as received, meso-lactide (PURAC) was purified by

distillation to remove traces of D- and L-lactide. The melting point of the purified meso-lactide was 54°C. Lactide mixtures were made up to the following ratios: 100 percent L-lactide, 90/10 L-lactide/meso-lactide, 5 70/30 L-lactide/meso-lactide, 50/50 L-lactide/meso-lactide, and 100 percent meso-lactide. Catalyst level was 2,500:1 molar ratio of initial monomer to tin with the tin being tin(II) bis (2-ethyl hexanoate) (Fascat[®] 9002). Lactic acid was added as a molecular weight 10 control agent to target a number average molecular weight of 50,000 (the same amount was added to all samples). Polymerization times were estimated to obtain conversions of 50 percent and 90 percent. For 120°C this was 4 hours and 16 hours, respectively. For 180°C 15 these times were 10 minutes and 50 minutes, respectively. Below in Table 15 are the GPC results (method of Example 8) of tests on the polymer samples produced by this procedure.

TABLE 15

	L/meso	Temp	Mn	Mw	PDI	%Conv
5	100% L	120°C	31,014 45,864	33,774 52,574	1.09 1.15	53.2 87.1
	100% L	180°C	27,785 56,839	32,432 98,125	1.17 1.73	46.7 93.3
10	90/10	120°C	34,541 29,222	38,586 34,466	1.12 1.18	62.3 89.3
	90/10	180°C	31,632 57,925	35,713 110,841	1.13 1.91	48.5 94.8
15	70/30	120°C	41,211 58,284	45,222 71,257	1.10 1.22	60.1 89.1
20	70/30	180°C	32,292 51,245	37,401 107,698	1.16 2.10	53.8 96.5
	50/50	120°C	15,888 25,539	17,969 31,834	1.13 1.25	57.8 90.6
25	50/50	180°C	34,375 44,590	42,018 98,028	1.22 2.20	62.5 95.5
	100% meso	120°C	33,571 45,237	40,635 68,142	1.21 1.51	73.4 94.3
30	100% meso	180°C	30,976 40,038	42,987 83,815	1.39 2.09	67.6 96.6

35

The results show that the ultimate number average molecular weight was not significantly affected by the temperature of polymerization, with an average of 41,000 at 120°C and 50,000 at 180°C. This implies that each lactic acid molecule initiates about one polymer chain, regardless of temperature. The ultimate weight average molecular weight is, however, significantly affected by temperature. At 120°C the weight average molecular weight averaged 52,000 and at 180°C the average was 100,000. This is believed to be due to a relative increase in the rate of transesterification at 180°C. The polydispersity index (PDI) at high conversion also reflects this, averaging 1.3 at 120°C and 2.0 at 180°C. It is believed these differences would have a

significant effect on the melt-processing characteristics of the polymer, with the higher weight average molecular weight of the polymer produced at 180°C expected to translate into better melt strength and processability.

These experiments show that polymerization at a higher temperature results in a polymer that is characteristically different. Further, the glass transition temperature for the samples polymerized at higher temperature is higher.

Example 14

Experiments with Stabilizing Agents and Metal Deactivators

15

Test 1

Conditions: vial polymerization, (Lactide is melted under a nitrogen-purged atmosphere in a round bottom flask with stirring. Catalyst and additives are added and aliquots of the mixtures are pipetted into silanized glass vials. Typically 5-10 grams of reaction mixture are used in a 16 ml. vial. The vials are tightly capped and placed into a preheated oil bath.) 10,000:1 molar ratio of lactide-to-tin, tin(II) bis(2-ethyl hexanoate) catalyst, 0.2 wt percent Ultrinox®626 in tetrahydrofuran (THF). 180°C. Time was 90 minutes.

The control with tin only polymerized to 84 percent conversion and reached a MWn of 31,700. The example with tin and Ultrinox® polymerized to 83 percent conversion and reached a number average molecular weight (MWn) of 39,800; an increase of 26 percent over the control.

The control sample turned light yellow, the sample with stabilizer remained colorless.

35

Test 2

Conditions: vial polymerization, 5000:1 molar ratio of lactide to tin, tin(II) bis(2-ethyl hexanoate) catalyst, 0.25 wt percent Ultrinox®626 (in THF). 180°C.

Time was 60 minutes. Lactide was used from the above described Gruber et al. process.

The control with tin alone polymerized to 67 percent conversion and reached a MWn of 62,900. The example
5 with tin and Ultrinox® polymerized to 66 percent conversion and reached a MWn of 75800; an increase of 21 percent over the control.

A second example with tin(II) bis(2-ethyl hexanoate), Ultrinox®, and 0.50 percent of Irganox® 1076,
10 which is a phenolic antioxidant, polymerized to 66 percent conversion and reached a number average molecular weight (MWn) of 74500; an increase of 18 percent over the control.

All samples were a dark yellow color, although the
15 samples with stabilizer had a slightly lower absorbance at 300 nm.

Test 3

Conditions: vial polymerization, 10,000:1 molar ratio of lactide to tin, tin(II) bis(2-ethyl hexanoate)
20 catalyst, 180°C, 80 percent L-lactide and 20 percent D,L-lactide purchased from Henley and Aldrich, respectively. Lactic acid was added to control molecular weight to about 75,000 at full conversion. One sample included 0.25 percent Ultrinox® 626 phosphite
25 stabilizer, one included 0.25 percent Irganox® 1076 antioxidant, and one control sample.

Samples were taken at various times and analyzed by GPC for conversion and molecular weight (the method of Example 8). The results are summarized in Table 16
30 below.

62

TABLE 16

	Time (hrs)	Control		Irganox®		Ultranox®	
		Mn	%conv	Mn	%conv	Mn	%conv
5	1	31,000	46	35,900	41	66,500	61
	2	45,400	74	56,800	74	102,700	83
10	4	69,600	93	74,100	93	97,200	91
	11	52,900	95	60,700	95	71,500	94

15

The sample with phosphite stabilizer polymerized faster, shown by the higher conversion at 1 and 2 hours, and went to a higher molecular weight than the control or the sample with Irganox®. The phosphite stabilized sample had a molecular weight more than 30 percent higher than the control for all time periods.

20

Test 4

The experiment above was repeated to compare the control to the phosphite-stabilized polymer, as summarized in Table 17 below.

25

TABLE 17

	Time (hrs)	Control		Ultranox®	
		Mn	%conv	Mn	%conv
30	1	36,600	37	71,500	59
	2	51,700	70	95,200	85
35	4	64,400	91	103,700	94
	8	58,100	96	95,700	94

40

The sample with phosphite stabilizer again polymerized faster and went to a higher molecular weight than the non-stabilized sample. The phosphite stabilized sample had a molecular weight more than 60% higher than the control for all time periods.

45

Test 5

Conditions: vial polymerization, 5,000:1 molar ratio of lactide to tin, tin(II) bis(2-ethyl hexanoate) catalyst, 180°C, 80 percent L-lactide and 20 percent D,L-lactide purchased from Henley and Aldrich. Lactic acid was added to control number average molecular weight to an estimated 80,000 at full conversion. One sample was run with 0.25 percent Ultrinox® 626 phosphite stabilizer, one with 0.25 percent Irganox® 1076 antioxidant, and one control sample.

Samples taken at various times and analyzed by GPC (the method of Example 8) for conversion and molecular weight. The results are tabulated in Table 18 below.

15

TABLE 18

	Time (hrs)	Control		Irganox®		Ultrinox®	
		Mn	Zconv	Mn	Zconv	Mn	Z conv
20	1	83,600	76	121,900	83	162,300	87
	4	74,400	93	104,300	95	123,900	86
	24	40,200	96	52,000	96	96,900	97
25	48	34,200	97	30,400	96	56,500	96
	72	25,000	96	22,400	96	69,500	96

30

The phosphite-stabilized sample had a molecular weight more than 60 percent higher than the control for all time periods. After 72 hours it had a molecular weight 2.8 times higher than the control. The sample with antioxidant showed an initial increase in molecular weight, relative to the control, but the effect disappeared after 48 hours.

The phosphite stabilized sample was significantly lighter in color than the control or the antioxidant treated sample.

40

Test 6

Conditions: vial polymerization, 5000:1 molar ratio of lactide to tin, tin(II) bis(2-ethyl hexanoate) catalyst, 0.25 wt percent Ultrinox®626 (in THF). 180°C.

- 5 Time was two hours. Gruber et al. process lactide washed with isopropyl alcohol was used.

The control with tin alone polymerized to 95 percent conversion and reached a number average molecular weight of 118,000. The example with tin and Ultrinox®
10 polymerized to 93 percent conversion and reached a number average molecular weight of 151,000, an increase of 28 percent over the control.

Test 7

- 15 Conditions: vial polymerization at 180°C. 5000:1 molar ratio of lactide to tin, tin(II) bis(2-ethyl hexanoate) catalyst. Lactide was 80 percent L-lactide and 20 percent D,L-lactide, purchased from Henley and from Aldrich. Lactic acid was added to target the
20 molecular weight to an Mn of 80,000. All stabilizers were added at 0.25 weight percent. Molecular weight (number average) was determined for samples pulled at 3 hours, while rate constants were based on samples pulled at 1 hour. The results of these screening tests on many
25 stabilizing agents following the above procedure are detailed below in Table 19. Product designations in Table 18 are tradenames or registered trademarks.

65

TABLE 19

	<u>Sample</u>	<u>MWn</u>	<u>% Conversion</u>	<u>Relative Rate</u>
5	Control 1	65,000	95.9	90
	Control 2	85,000	95.9	100
	Control 3	76,000	96.6	100
	Control 4	69,000	96.2	100
10	Control 5	74,000	96.8	110
	Control 6	70,000	97.2	110
<u>PHOSPHITES</u>				
15	Ultranox 626 (GE)	103,000	96.8	100
	Weston TDP (GE)	64,000	70.0	60
	Weston PDDP (GE)	67,000	76.7	60
	Weston PNPG (GE)	92,000	94.1	100
	Irgafos 168			
20	(Ciba-Geigy)	95,000	95.3	120
	Weston 618 (GE)	99,000	95.1	100
	Sandostab P-EPQ (Sandoz)	108,000	94.7	110
	Weston TNPP (GE)	88,000	97.9	130
25				
<u>PHENOLIC ANTIOXIDANTS</u>				
	Irganox 1010			
30	(Ciba-Geigy)	95,000	97.5	110
	Cyanox 1790 (Cyanamid)	98,000	96.9	120
	BHT	87,000	96.5	130
	Irganox 1076			
35	(Ciba-Geigy)	121,000	97.8	130
	Topanol CA (ICI)	84,000	96.6	160
<u>AMINES</u>				
40	Tinuvin 123 (Ciba-Geigy)	65,000	94.8	70
	Tinuvin 622 (Ciba-Geigy)	82,000	95.7	80
45	Naugard 445 (Uniroyal)	93,000	98.2	120
<u>THIOETHER</u>				
50	Mark 2140 (Witco)	77,000	97.0	120

METAL DEACTIVATORS

5	Irganox MD1024 (Ciba-Geigy)	34,000	65.7	10
	Naugard XL-1 (Uniroyal)	91,000	95.8	110

10

Note, that with a few exceptions, the phosphites and the phenolic antioxidants provide increased molecular weight with no reduction in polymerization rate. Of the amines, only Naugard® 445 provided stabilization without a rate decrease. The metal deactivators are expected to deactivate the catalyst, as was observed for Irganox® MD1024. The Naugard® XL-1 did not accomplish deactivation.

20

Example 15Polymer Melt Stability as a Function of Moisture Content

Lactide, produced and purified in a continuous (Gruber et al.) process, was fed at a rate of 3 kg/hr to a continuous polymerization pilot plant. Catalyst was added with a metering pump at the rate of 1 part catalyst to 5000 parts lactide on a molar basis. The reaction system was blanketed with nitrogen. The reactor vessels consist of two continuous stirred tank reactors (CSTR) in series. The first had a 1-gallon capacity and the second had a 5-gallon capacity. The reactors were run 60-80 percent liquid filled and at 170-180°C. Polymer melt pumps moved the liquid from CSTR 1 to CSTR 2, and from CSTR 2 through a die into a cooling water trough. The polymer strand thus produced was pulled from the trough by a pelletizer and stored as pellets.

The pelletized poly(lactide) was put into a drying hopper and dried at 40°C under flowing dry air. Samples were pulled after one hour and four hours. These samples were then run through a single screw Brabender® extruder, with a retention time of approximately 3

minutes. Samples were analyzed for moisture by an automatic Karl Fischer apparatus and for molecular weight by GPC (the method of Example 8). The results of these tests are documented in Table 20 below.

5

TABLE 20

	<u>Sample</u>	<u>Extruder Temperature (C)</u>	<u>Weight Average Molecular Weight</u>
10	Initial		63,000
	Dried 1 hour (1200 ppm H ₂ O)	137	44,000
15		145	48,000
		162	35,000
		179	30,000
	Dried 4 hours (150 ppm H ₂ O)	140	63,000
20		140	69,000
		160	65,000
		178	68,000

25 These results show the detrimental effect of water in the lactide polymer resin during melt polymerization and the need to properly dry the poly(lactide) before melt-processing.

30

Example 16Degradation of Crystalline and
Amorphous Poly(lactide)

Two literature references disclose poly(D,L-lactide) to degrade faster than poly(L-lactide), attributing the result to crystallinity of poly(L-lactide). These are: Kulkarni et al., J. Biomed. Mater. Res., vol. 5, pp. 169-181, (1971); Makino et al., Chem. Pharm. Bull., vol. 33, pp. 1195-1201, (1985). An experiment was conducted to measure the effect of crystallinity on polymer degradation and is detailed below.

An amorphous poly(lactide) sample (clear, and less than 1 percent crystallinity based on DSC) and a

crystalline poly(lactide) sample (opaque, and approximately 50 percent crystallinity based on DSC) were subjected to biodegradation in a compost test (50°C, with aeration). The DSC apparatus was a TA
5 Instruments, Inc., model 910 differential scanning calorimeter with DuPont 9900 computer support system typically programmed to heating at a rate of 10°C per minute to 200°C. The samples had different optical composition, with the crystalline sample being more than
10 90 percent poly(L-lactide) and the amorphous sample being less than 80 percent poly(L-lactide) with the balance being either poly(D,L-lactide) or poly(meso-lactide). Samples of each polymer were subjected to a compost test (ASTM D 5338) which included mixing a
15 stabilized compost and providing a source of humidified air while maintaining a temperature of about 50°C. The amorphous sample was completely degraded after 30 days of composting. The crystalline sample was only 23 percent degraded based on carbon dioxide after the same
20 period of time.

Additional samples of these two polymers were subjected to chemical hydrolysis at 50°C (hydrolysis is believed to be the rate-limiting step in the biodegradation process). The chemical hydrolysis
25 procedure included placing 0.1 gram poly(lactide) in 100 ml of 0.2M phosphate buffer (pH = 7.4). The samples were held for 1 week, then filtered, washed with deionized water, and dried at 25°C under vacuum. The initial weight average molecular weight for each sample
30 was about 70,000. After 1 week the amorphous sample had a weight average molecular weight of 10,000 and the crystalline sample had a weight average molecular weight of 45,000, determined by GPC (the method of Example 8). Neither sample had significant weight loss at this time.
35 Both of these tests demonstrate that degradation of crystalline poly(lactide) is slower than degradation of amorphous poly(lactide).

Example 17Effect of Monomer Concentration on Film Modulus

Poly(lactide) was precipitated in methanol from a
 5 chloroform solution in order to remove the residual
 lactide monomer. GPC analysis (the method of Example 8)
 showed the precipitated polymer to contain 0.0 percent
 lactide.

The polymer was dissolved in chloroform to make a 10
 10 wt percent solution, and lactide was added back to make
 5 separate solutions which, after removing the
 chloroform, are calculated to produce films containing
 0.0, 0.2, 0.4, 1.0 and 4.0 weight percent lactide in
 poly(lactide). These solutions were solvent cast onto
 15 glass, dried overnight at room temperature in a fume
 hood, and removed to a vacuum oven. The films were hung
 in the vacuum oven and dried at 30°C for 72 hours. GPC
 analysis of the vacuum-dried films showed measured
 lactide levels of 0.0, 0.0, 0.4, 0.7 and 3.7 wt percent.

20 The films were then tested for film modulus using
 ASTM procedure D882.

The results are shown below in Table 21.

TABLE 21

	<u>Z</u>	<u>Tensile</u>	<u>Std.</u>	<u>Z</u>	<u>Std.</u>	<u>Elastic</u>	
	<u>Lactide</u>	<u>(psi avg.)</u>	<u>Dev.</u>	<u>Elongation</u>	<u>Dev.</u>	<u>Modulus</u>	<u>Std.</u>
						<u>(psi avg.)</u>	<u>Dev.</u>
30	0	5490	636	2.85	0.14	730,000	103,000
	0	6070	123	2.85	0.22	818,000	35,000
	0.4	5670	227	2.75	0.27	779,000	44,000
	0.7	5690	343	4.04	1.12	749,000	58,000
	3.7	5570	458	3.33	1.43	738,000	66,000

35

Example 18Rate of Water Uptake Versus Optical Composition

Samples of poly(lactide), made from 80 percent L-
 lactide and 20 percent of either D,L-lactide or meso-
 40 lactide, were ground to pass a 20 mesh screen. The
 samples were dried and devolatilized under vacuum then

removed to a constant humidity chamber maintained at 24°C and 50 percent relative humidity. The rate of moisture pick-up was determined gravimetrically, with the final results verified by Karl-Fischer water analysis. The rate of moisture pickup is shown below in Table 22.

TABLE 22

Time (Minutes)	Parts Per Million Weight Gain	
	<u>L/D,L Polymer</u>	<u>L/Meso Polymer</u>
10	600	1000
15	1100	1500
60	1500	1800
120	1600	2100
870	2100	2600
Final (Karl-Fischer)	3000	2600

Example 19Standard Test of Melt Stability

A standard test for determining melt stability is as follows:

A small sample (200 grams or less) of polymer is ground or pelletized and devolatilized by holding under vacuum (about 10 mm Hg) at a temperature of 130°C or less for 18 hours. At this point the residual lactide content should be 1 wt percent or less. Portions (1-5 grams) of the devolatilized sample are then placed in a 16 ml sample vial, tightly capped, and placed in a 180°C oil bath. Samples are removed at times of 15 minutes and 1 hour and analyzed for lactide content by GPC or other appropriate techniques. Lactide which may collect on the cooler portions of the vial is included in the product work-up and test.

Melt-stabilized poly(lactide) will show less than 2 percent lactide in the 15 minute sample, and more preferably less than 2 percent lactide in the 1 hour sample. The most highly stabilized poly(lactide)s will maintain lactide contents of less than 1 percent in both

the 15 minute and 1 hour samples, preferably less than 0.5 percent. An unstabilized poly(lactide) may reach the equilibrium lactide content at 180°C of 3.6 wt percent, or may go even higher as lactide is driven from the polymer melt and collects on the cooler top walls of the vial.

Example 20

Water Scavenger Experiments

10 Dried poly(lactide) pellets were processed in a twin screw extruder to devolatilize and to prepare a portion with 0.5 percent by weight of a water scavenger (Stabaxol® P). The strands leaving the extruder are cooled in a water trough and chopped into pellets.

15 Samples of the control and the test sample were then analyzed by the Karl Fischer technique for moisture content, with no drying. The control sample contained 1700 ppm water, the test sample had 450 ppm water. The control sample was then dried under nitrogen at 40°C,

20 reducing the water content to 306 ppm. A vacuum-dried control sample had 700 ppm water.

The as-produced test sample and the dried control samples were then processed in a 1/2" single screw extruder (Brabender®) at 160°C, with a retention time of

25 3 minutes. The number average molecular weight for the dried control sample dropped from an initial value of 44,000 to a final value of 33,000 for the 306 ppm water sample and to 28,000 for the 700 ppm water sample. The test sample number average molecular weight dropped from

30 an initial value of 40,000 to a final value of 33,000.

This sample shows how the water scavenger protected the polymer from moisture pick-up, imparting the same stability as a thorough drying of the control sample. Combining a water scavenger with appropriate drying is

35 expected to give even greater stability.

Example 21Optimization of Catalyst Concentration

A mixture of 80 percent L-lactide and 20 percent D,L-lactide was polymerized using three different levels of tin(II) bis(2-ethyl hexanoate) catalyst. Batches were prepared at initial monomer/catalyst molar ratios of 1000:1, 3000:1, and 20,000:1. Polymerization times were adjusted to reach high conversion without being excessively long and thereby causing degradation in the melt. The reaction times were 1, 2 and 20 hours, respectively. The polymerization temperature was 180°C. The polymers were ground to a coarse powder and devolatilized at 125°C and 10 mm Hg overnight. The samples were then reground and 1-gram portions of each were placed into silanized vials, 16 ml capacity. The vials were sealed and placed into an oil bath at 180°C. Vials were then removed at various times and the samples were analyzed by GPC after dissolution in chloroform. The molecular weights and lactide contents are shown below in Table 23.

TABLE 23

	Sample	Time (min)	Number Average Molecular Weight	Weight Average Molecular Weight	Lactide Weight %
25	1000:1	0	39,000	81,300	0.8
		5	28,100	57,300	2.4
		15	25,800	49,700	2.8
		30	23,100	43,800	3.7
		60	22,800	43,200	3.6
35	3000:1	0	53,100	113,600	0.6
		5	39,000	76,400	0.4
		15	30,300	65,400	1.9
		30	29,000	60,400	2.7
		60	28,200	55,200	2.8
40	20000:1	0	89,200	184,000	0.0
		5	81,200	165,100	0.0
		15	54,300	134,600	0.1
		30	51,100	119,600	0.0
		60	49,500	111,000	0.0

These results show the benefit of optimizing the catalyst level used in the polymerization process. Note that both lactide reformation and molecular weight retention benefits are realized from the reduced catalyst levels (higher monomer/catalyst ratio).

It is believed catalyst levels should be limited to 1000:1 for the high end of catalyst usage, with 3000:1 being more preferable and showing somewhat improved stability. Lower levels still, such as 20000:1, show greatly improved stability. Beyond this level it is believed the polymerization rates become too slow to be practical.

Example 22

Removal of Tin Catalyst from Poly(lactide) by Precipitation

Forty-five grams of L-lactide and 13 grams of D,L-lactide were charged with 78 milligrams of crystalline lactic acid to a 200 ml round bottom flask. This was heated to 180°C with magnetic stirring in an oil bath and blanketed with dry nitrogen. Catalyst in the form of tin(II) bis(2-ethyl hexanoate) was added as 0.20 ml of a 0.47 g/ml solution in THF after the molten lactide was at temperature. The mixture was allowed to stir for one minute and then pipetted into 3 silanized glass vials, which were then sealed and placed into a 180°C oil bath for 75 minutes. The vials were allowed to cool and the polymer recovered by breaking the glass. The polymer was ground to a coarse powder and dissolved in chloroform to make a 10 percent solution. The polymer contained 3.8 percent residual monomer and had a number average molecular weight of 70,000 as determined by GPC measurement (the method of Example 8).

Five hundred milliliters of methanol were placed in a 1-liter glass blender flask. The blender was turned on to medium speed and 50 ml of the polymer in chloroform solution was poured in over a period of three minutes. After one additional minute of blending the

mixture was filtered, then rinsed with 100 ml of methanol, and dried overnight under vacuum. The polymer consisted of a fibrous mat. It contained 0.3 percent residual monomer and had a number average molecular weight of 66,900.

The measured tin level in the precipitated polymer was 337 ppm by weight, compared to a calculated value of 466 ppm for the as-produced polymer. This result indicates the feasibility of reducing residual catalyst levels in lactide polymers by solvent precipitation with the benefit of improved stability as detailed in Example 21.

Example 23

15

Shear Rate

Samples of devolatilized poly(lactide) were tested in a Rosand Model 14°C capillary rheometer. The die was 1 mm diameter and 16 mm long, with an entry angle of 180 degrees. The table below gives the pressure drop across the die as a function of nominal shear rate (not Rabinowitsch corrected) for various molecular weights and temperatures.

75

Table 24Results at 150°C.

5			Temp. (°C)	Nominal shear rate	Pressure Drop (MPa)
	<u>Mn</u>	<u>MW</u>		(s ⁻¹)	
10	34,000	70,000	150	192	2.0
				384	5.5
				960	10.0
				1920	13.8
				4800	19.7
				9600	23.7
15	52,000	108,000	150	192	9.9
				384	15.6
				960	19.9
				1920	23.9
				4800	29.4
20				9600	----
25	60,000	137,000	150	192	7.4
				384	11.1
				960	16.6
				1920	21.0
				4800	----
				9600	----
30	183,000	475,000	150	192	19.1
				384	27.0
				960	31.4
				1920	----
				4800	----
35				9600	----

<u>Results at 175°C.</u>					
5	Pressure		<u>Temp. (°C)</u>	Nominal	<u>Drop (MPa)</u>
	<u>Mn</u>	<u>MW</u>		<u>shear rate (s⁻¹)</u>	
10	34,000	70,000	175	192	0.4
				384	0.5
				960	3.4
				1920	5.5
				4800	9.2
				9600	12.5
15	52,000	108,000	175	192	2.2
				384	4.6
				960	7.6
				1920	11.5
				4800	17.2
20				9600	22.1
25	183,000	475,000	175	192	11.5
				384	16.6
				960	20.2
				1920	24.4
				4800	29.9
				9600	----
30					
<u>Results at 200°C.</u>					
35	Pressure		<u>Temp. (°C)</u>	Nominal	<u>Drop (MPa)</u>
	<u>Mn</u>	<u>MW</u>		<u>shear rate (s⁻¹)</u>	
40	60,000	137,000	200	192	0.5
				384	1.6
				960	3.3
				1920	5.3
				4800	----
				9600	13.2
45					
50	183,000	475,000	200	192	7.0
				384	11.0
				960	14.2
				1920	17.9
				4800	21.6
				9600	----

Example 24Effect of Meso-lactide Concentration
on Rate of Crystallization.

5 Polymer samples of various optical compositions were prepared by polymerizing mixtures of L-lactide and meso-lactide with tin(II)bis(2-ethyl hexanoate) catalyst at a temperature of about 180°C. A portion of each sample was tested in a Mettler Differential Scanning

10 Calorimeter (DSC), Model 30, by heating from 60°C to 200°C at 20° C/minute. The sample was then held at 200°C for 2 minutes to completely melt any crystals. The samples were then quenched to the annealing temperature of interest and held 15 minutes. The

15 samples were then quenched to 60°C and reheated at 20°C/minute to 200°C to determine the crystallinity. The crystallinity of the sample following annealing is proportional to the energy of the melting endotherm minus any crystallization exotherm.

20

Table 25

Net endotherm following 15 minute annealing
at given temperature (J/gm)

25	Sample (% meso)	Temperature = 85°C	Temperature = 110°C
	0	34.3	48.4
	3	5.1	48.2
30	6	0.1	14.5
	9	0.3	11.0

The results show that introducing meso-lactide greatly reduces the rate of crystallization for

35 poly(lactide). Therefore, control of the meso level and tailoring the processing conditions are important.

Example 25Process Stabilizer Effect on Rate of Hydrolysis

40 Poly(lactide) was produced using a vial polymerization of lactide at 180°C for 2 hours. The

catalyst (tin(II) bis(2-ethyl hexanoate)) was used in a molar ratio of 1 part catalyst to 10,000 parts lactide. One of the samples had 0.25 wt-% of the process stabilizer PNPG (commercially available from General

5 Electric) added to the molten lactide.

Hydrolysis was carried out in 0.2 M phosphate buffer (pH 7.4) at 75°C. Results with and without the stabilizer are shown in the table below:

10

Table 26

Day	<u>Weight Average Mol. Wt.</u>		<u>Weight % Remaining</u>	
	<u>With PNPG</u>	<u>Without</u>	<u>With PNPG</u>	<u>Without</u>
0	127,000	116,000	100	100
15 1	23,700	24,500	84	82
2	4,600	10,700	70	74
3	12,600	6,000	63	64
4	5,700	4,800	74	68
5	4,500	3,800	75	79
20 7	3,800	4,600	42	62

The results show that the process stabilizer had no effect on the rate of hydrolysis.

25

Example 26

The Effect of Meso-lactide Concentration
on Crystallization

Samples of devolatilized poly(lactide) of varying optical composition and with number average molecular weights in the range of 50,000 to 130,000 were prepared in a continuous pilot plant. The samples were dissolved in chloroform to a concentration of 5 grams/100cc and the optical rotation of the samples was measured to determine the concentration of meso-lactide which had been present in the monomer mixture prior to polymerization. Separate optical rotation and gas chromatography analysis of the monomer mixture confirmed that L-lactide and meso-lactide are the predominate components when meso-lactide is present at a

concentration of 20 percent or less, and only a small correction is required for D-lactide.

Additional samples were made by polymerizing mixtures with known weights of L-lactide and meso-
5 lactide.

All samples were subjected to an annealing procedure to develop crystallinity. The annealing procedure consisted of placing the samples in an oven at 100-105°C for 90 minutes, then lowering the temperature 10°C each
10 1/2 hour until the temperature reached 45°C. The oven was then shut off and the samples were allowed to cool to room temperature. The energy of the melting endotherm and the peak melting temperature were then measured using a Mettler Differential Scanning
15 Calorimeter (DSC) apparatus with a scan speed of 20°C/minute. The energy of melting is a measure of crystallinity in the annealed samples.

Figure 3 shows sharp decline in crystallinity between 9 and 12 percent meso content.
20

Example 27

Effect of Plasticizer on Crystallization Rate

Devolatilized polymer samples from a continuous pilot plant were compounded with dioctyl adipate (a
25 plasticizing agent) and/or silica in a twin screw extruder. The samples were then tested from crystallization rate using the DSC of Example 24. In this case the DSC program included a first upheal, in which the samples were heated from -20°C to 200°C at a
30 rate of 20°C/minute, holding at 200°C for 2 minutes, quenching, and a second upheal from -20°C to 200°C at 20°C/minute. The energy of the crystallization exotherm, occurring at a temperature from about 75°C to about 115°C, is proportional to the amount of
35 crystallization which occurs during this two minute period.

The table below shows the increased crystallization observed when the plasticizer dioctyl adipate (DOA) is present, either with or without silica present. The base polymer, without plasticizer, shows no
 5 crystallization during the DSC upheal. The exotherms are reported on a joules per gram of poly(lactide) basis (filler free basis).

10

Table 27

	<u>Sample</u>	<u>First Upheat Exotherm (J/gm)</u>	<u>Second Upheat Exotherm (J/gm)</u>
15	Base polymer	0	0
	Base polymer+ 8 wt% DOA	26.7	27.2
20	Base polymer+ 40 wt% silica	4.0	0
25	Base polymer+ 40 wt% silica+ 5 wt% DOA	27.1	27.6

Example 28An Evaluation of Nucleating Agents

30 A devolatilized sample of poly(lactide) polymer was compounded with a variety of potential nucleating agents in a single screw extruder. The candidate nucleating agents were added at a nominal level of 5 percent by weight. The single screw extruder is not as effective
 35 of a mixer as would be used commercially, so failure to observe an effect in these tests does not mean that a candidate agent would not be effective if blended more thoroughly. A positive result in this test demonstrates potential ability to increase crystallization rates.

40 Additives which increased crystallinity in the second upheal (on a quenched sample) were rated ++, additives which showed an effect only on the first upheal were rated +, and additives which showed no effect were rated 0.

Table 28

	<u>Additive</u>	<u>Effect</u>
5	None	0
	talc, MP1250 (Pfizer)	++
	3-nitro benzoic acid	0
	saccharin, sodium salt	++
	terephthalic acid, disodium salt	0
10	calcium silicate, -200 mesh	+
	sodium benzoate	+
	calcium titanate, -325 mesh	+
	boron nitride	+
	calcium carbonate, 0.7 micron	0
15	copper phthalocyanine	+
	saccharin	0
	low molecular weight polyethylene	0
	talc, Microtuff-F (Pfizer)	++
	talc, Ultratalc (Pfizer)	++
20	ethylene acrylic acid sodium ionomer (Allied Signal)	0
	isotactic polypropylene	+
	polyethylene terephthalate	0
	crystalline poly(L-lactide)(low mol. wt.)	++
25	Millad 3940 (Milliken)	++
	Millad 3905 (Milliken)	+
	NC-4 (Mitsui)	+
	polybutylene terephthalate	+
	talc in polystyrene (Polycom Huntsman)	+
30	talc in polyethylene (Advanced Compounding)	++

35

Example 29Heat Set Crystallization of an Oriented Poly(lactide) Film

Two film samples, one non-oriented and the other biaxially oriented, were constrained in a film holder and annealed for either 5 minutes or 15 minutes in an oil bath at 85°C. The extent of crystallization was determined by DSC from the melting endotherm of the crystalline domains formed during the annealing, using a ramp rate of 20°C/minute. The biaxially oriented film developed significantly more crystallinity for each time, as shown in the table below.

45

Table 29

<u>Sample</u>	<u>t=0 minutes endotherm(J/gm)</u>	<u>t=5 minutes endotherm(J/gm)</u>	<u>t=15 minutes endotherm(J/gm)</u>
5 non-oriented	0.6	0.0	0.8
10 biaxially oriented	0.7	8.1	8.5

Each of the films was made from lactide mixtures containing an estimated meso-lactide content of about 12 wt%, with about 88 wt% L-lactide. When subjected to the slow oven annealing procedure of Example 26, samples from both of the films developed crystallinity which gave a melting endotherm of about 25 J/gm. The biaxially oriented film had been stretched approximately 4x in the machine direction and 2x in the transverse direction (using a tenter frame), all at about 63-74°C.

Example 30Properties of PLA-cellulose Acetate Blends

In the twin screw extruder described in Example 2, poly(lactide) was blended with cellulose acetate (Tenite 110 from Eastman), cellulose acetate propionate (Tenite 375 from Eastman), and cellulose butyrate (Tenite 575 from Eastman) in levels shown in Table 30. The poly(lactide) had a weight average molecular weight of about 200,000, a meso content of about 10.5%, a residual lactide level of about 0.3%, and a moisture content less than about 300 ppm. The weight percentages of poly(lactide) and cellulose derivatives is shown in the following table:

35

Table 30

Composition	1	2	3	4	5	6	7	8	9	10
CA		20%	50%	80%						
CAP					20%	50%	80%			
5 CAB									20%	50%
80%										
PLA	100%	80%	50%	20%	80%	50%	20%	80%	50%	20%

- The blends were prepared using a melt temperature of 190-200°C, die pressure of 18-20 bar, and a screw speed of 250 rpm. All materials but one were pelletized, dried and injection molded into a standard ASTM test specimen mold for testing. Composition 3 was not pelletized due to poor strength of the extruded polymer stand. This is evidence of mechanical incompatibility of the cellulose acetate and poly(lactide). The remaining compositions were tested for tensile strength and elongation to break in accord with ASTM D 638-91 and heat distortion temperature according to ASTM D 648-82.
- Another test performed on each sample was a test for resistance to hot water. A standard tensile bar is placed in a hot water bath at 190°F for 3 minutes. The sample is removed from the water and placed in a horizontal position with the flat sides facing up and down with one end of the tensile bar held in a clamp. Samples which have softened as a result of the hot water exposure will bend under the force of gravity. The degree of bend is measured using a protractor. The maximum degree of bend is 90° and constitutes a failure.
- The best result is no bend of the sample.

The test results are shown in the following table:

Table 31

<u>Sample</u>	<u>Tensile strength to break (psi)</u>	<u>Elongation to break (%)</u>	<u>Heat distortion temperature (°C)</u>	<u>Hot water test angle of (deformation)</u>
1	7100	5.4	51.1	46°
2	5900	1.5	41.7	44°
3	-----	-----	-----	-----
4 10	3200	2.2	-----	0°
5	8100	2.2	46.7	10°
6	5600	6.9	45.7	3°
7	4700	17.9	50.0	0°
8	7600	2.3	48.2	23°
9 15	5500	8.3	-----	0°
10	5100	17.9	51.9	0°

These results show that many compositions of poly(lactide) blended with CA, CAP and CAB are compatible enough to exhibit good physical properties. Blends of PLA with high loadings of CAP and CAB may
5 increase the flexibility of unblended poly(lactide). Resistance of poly(lactide) to hot water was dramatically increased by addition of CA, CAP or CAB.

Example 31

10 Effect of Talc on Crystallization Rate of Poly(lactide)

Poly(lactide) prepared in a continuous pilot plant, from a lactide mixture with an approximate composition of 91% L-lactide and 9% meso-lactide, was dried and devolatilized in a twin screw extruder. The polymer was
15 then redried and talc (Ultratalc 609, from Pfizer) was compounded in at levels from 2 wt% to 21 wt%.

Samples of the compounded poly(lactide) were placed in a DSC apparatus and subjected to a heating program consisting of a first upheal from 60°C to 200°C for two
20 minutes, quenching to 90°C, holding at 90°C for 15 minutes, followed by a quench and second upheal from 60°C to 200°C. The first upheal and quench is to make each of the samples amorphous, the crystallization exotherm is measured during the 90°C isothermal run, and
25 the second upheal is used to confirm the isothermal exotherm through a direct measurement of the melting endotherm of the crystalline domains formed during the isothermal annealing.

The table below shows the extent of crystallization
30 (reported on talc free basis) as a function of time for various talc loadings. The results show that talc significantly increases the rate of crystallization of the poly(lactide).

Table 32

<u>Sample</u>	<u>Exotherm (J/gm of PLA) to time</u>					<u>2nd upheat</u> <u>J/gm</u>
	<u>2 min.</u>	<u>5 min.</u>	<u>8 min.</u>	<u>12 min.</u>	<u>15 min.</u>	
PLA	0	0	0	0	0	0
PLA+ 2 wt% talc	0.1	0.6	1.1	4.2	7.9	10.0
PLA+ 6 wt% talc	0.5	2.6	5.3	11.8	17.2	21.2
PLA+11 wt% talc	0.8	2.5	10.8	22.2	27.3	28.2
PLA+21 wt% talc	2.3	11.3	19.2	25.4	27.6	26.2

Example 32Stress-induced Crystallization
of a Poly(lactide) Sheet

Poly(lactide), copolymerized with 0.55 wt%
 5 epoxidized linseed oil, from a continuous pilot plant
 was dried, devolatilized, redried, and compounded with
 25 wt% talc (Ultratalc 609, Pfizer) and 25 wt%
 plasticizer (Citroflex A-4, Morflex). This material was
 then dried and cast into a 15 mil sheet. Strips of the
 10 sheet, approximate 1"x4", were loaded into an MTS test
 instrument using a 1" gap width and stretched to about
 4x at room temperature. After stretching, the test
 samples were removed from the MTS and analyzed by a DSC
 to determine the amount of stress-induced
 15 crystallization, using the net melting endotherm.
 Crystallization exotherms were not observed for these
 samples, except for the unstretched control.

Table 33

20	<u>Sample (stretching speed)</u>	<u>Melting endotherm (J/gm)</u>
	unstretched (control)	2.4 (net, after subtracting crystallization exotherm)
25	5 mm/sec (nominal)	19.1
	26 mm/sec (estimated)	17.3
30	26 mm/sec (estimated)	19.2
	26 mm/sec (estimated)	17.5
35	26 mm/sec (estimated)	18.3 (sample tore at end of test)

Each of the stretched samples showed a pronounced
 development of stress-included crystallization.

40

Example 33Stress-induced Crystallization of a Poly(lactide) Film

Poly(lactide), copolymerized with 0.55 wt%
 epoxidized linseed oil, was dried, devolatilized, and

redried. The lactide had a meso-lactide content of about 8-10%. The material was then cast into 15 mil sheet. Squares of the material, approximately 5"x5", were stretched using an Iwamoto biaxial stretcher either in a single direction or biaxially. The stretching temperature was 65°C and the stretching speed was 10 mm/s. The stretched films were tested by DSC to determine the extent of crystallization.

10

Table 34

	Sample	Cold Crystallization Exotherm (J/gm)	Melting Endotherm (J/gm)	Net J/gm
15	1x2	1.2	1.7	0.5
	1x3	16.4	23.7	7.3
	1x4	12.6	19.9	7.3
20	4x4	17.7	22.6	4.9

Example 3425 Crystallization of Pellets and Films During Storage

The talc filled, plasticized, pellets and unoriented films of Example 32 were stored at room temperature for 12 days and 11 days, respectively. They were then retested by DSC to determine if any crystallization had taken place during storage. Significant crystallization had occurred, and the feed pellets showed a melting endotherm of 17.4 J/gm and the film showed a melting endotherm of 18.8 J/gm on an as-tested basis. This corresponds to 35 J/gm and 38 J/gm on a poly(lactide) basis. No crystallization exotherms were observed during the DSC upheat.

Example 3540 Stress-induced Crystallization and Strain
Hardening of a Poly(lactide) Film

Additional samples of the polymer film used in Example 33 were subjected to uniaxial stretching on the

Iwamoto biaxial stretcher. The polymer, after melting and quenching, exhibited a T_g with an inflection point of 59°C when tested by DSC at a scan rate of 20°C/min. The samples were stretched at various temperatures at a rate of 99 mm/sec and tested by DSC to determine the extent of crystallization. Force curves for the stretching operation were also recorded.

res

10
10

Table 35

5	Stretching Temperature (°C)	Net Melting Endotherm (J/gm)	Stretching Force (kg) needed at elongation of <u>initial</u>	<u>200%</u>	<u>300%</u>	<u>400%</u>
	68	18	28	24	27	49
	72	18	26	30	40	52
10	76	20	14	14	14	16
	80	8	11	11	12	16

The table above shows the development of stress induced crystallinity when the polymer was stretched at temperatures up to 20°C above T_g . At 80°C less stress induced crystallinity developed, although the sample
5 crystallized readily during subsequent annealing. The samples stretched at 68°C and 72°C show pronounced strain hardening at 300-400% elongation during this test. The samples stretched at higher temperatures did not show the same degree of strain hardening.

10 It will be understood that even though these numerous characteristics and advantages of the invention have been set forth in the foregoing description, together with details of the structure and function of the invention, the disclosure is illustrative only, and
15 changes may be made in detail, especially in matters of shape, size and arrangement of the parts or in the sequence or the timing of the steps, within the broad principle of the present invention to the full extent indicated by the broad general meaning of the terms in
20 which the appended claims are expressed.

These results show that many compositions of poly(lactide) blended with CA, CAP and CAB are compatible enough to exhibit good physical properties. Blends of PLA with high loadings of CAP and CAB may
25 increase the flexibility of unblended poly(lactide). Resistance of poly(lactide) to hot water was dramatically increased by addition of CA, CAP or CAB.

Example 36

30 Effect of Optical Composition on Hydrolysis Rate

Poly(lactide) samples of various optical composition were prepared by melting, under nitrogen, the lactide monomers and adding catalyst to the molten monomer. Catalyst (tin (II) bis 2-ethyl hexanoate) was added on a
35 molar basis in the ratio of 1 part catalyst to 500 parts lactide. The mixture was placed in a sealed vial and

polymerized in a hot oil bath at a temperature of 180°C for 2 hours. Conversion to polymer was 92% or higher.

The hydrolysis test consists of placing 0.1 gram of polymer pellets into 100 ml of a 0.2 phosphate buffer solution with pH = 7.4. The solution is maintained at 75°C through the test. A series of samples are made up and pulled at various times throughout the test. The contents are collected on a #2 Whatman filter, washed with deionized water, and dried at 25°C under vacuum.

The table below shows the weight average molecular weight (determined by GPC in chloroform at 35°C, calibrated against polystyrene standards) and the percent residue remaining from samples made from 100% L-lactide, 100% D,L-lactide, and 100% meso-lactide. Each of the samples was initially amorphous. The sample made from L-lactide showed potential crystallinity of 47 J/gm when subjected to the standard annealing procedure of Example 26. The samples made from D,L-lactide and from meso-lactide showed endotherms of 1.2 and 0 J/gm after annealing, respectively, thus indicating non-crystallizing samples.

Table 36

Day	Weight Average Mol. Wt.			Weight % Remaining		
	L	D,L	meso	L	D,L	meso
0	44,600	36,800	42,400	100	100	100
1	38,800	14,700	11,900	76	77	81
2	20,000	8,100	7,200	69	43	24
3	11,000	5,400	5,000	73	12	3
4	16,500	4,300	5,000	84	14	9
5	7,300	4,300	5,000	82	19	10
7	7,000	4,300	5,000	74	7	7

The results show that poly(lactide) made from pure L-lactide hydrolyzes much slower than does poly(lactide) made from either D,L-lactide or meso-lactide. This is believed to be due to crystallization of the poly(L-lactide) during the hydrolysis test.

Example 37Effect of Annealing on Hydrolysis of
Samples with the Same Optical Composition

Poly(lactide) pellets, produced in a continuous pilot plant and then dried, devolatilized, and dried again, were subjected to the hydrolysis test of Example 36. The optical composition of the lactide in the pilot plant is estimated to be about 90 wt% L-lactide and about 10 wt-% meso-lactide. One of the samples was subjected to the standard annealing process prior to the hydrolysis procedure, developing 37 J/gm of crystallinity. The second sample was not annealed, and was initially amorphous. (less than 1 J/gm) The table below shows the hydrolysis results for the two samples as a function of time.

Table 37

Day	Weight Average Mol. Wt.		Weight % Remaining	
	Annealed	Not annealed	Annealed	Not annealed
20				
0	103,000	106,700	100	100
2	38,900	20,900	84	81
3	15,500	12,400	58	52
6	12,700	10,000	55	38
25				
30	11,900	11,000	33	30
60	9,200	9,000	25	23
90	9,200	9,000	10	9
120	9,200	9,000	0	0

The results show that both the annealed sample and the amorphous sample hydrolyzed at the same rate. This is believed to be due to crystallization of the initially amorphous sample during the hydrolysis test.

Example 38Demonstration of Crystallization during Hydrolysis

Poly(lactide) was dried, devolatilized, redried, and compounded with 15% talc and 2% TiO₂. The lactide used to make the polymer was predominately L-lactide, with about 11% meso-lactide. The pellets were dried again and cast into a sheet suitable for thermoforming.

Samples of two thermoformed articles were then subjected to a hydrolysis test in 0.2 M phosphate buffer. The temperature of the solution was adjusted following the ASTM compost profile (ASTM D 5338), consisting of day 1 at 35°C, days 2-5 at 58°C, days 6-27 at 50°C, and days 28-45 at 35°C. The results of the test are shown below:

Table 38

	Day	Weight Average Mol. Wt.		Crystallinity (J/gm)	
		Article 1	Article 2	Article 1	Article 2
10	0	230,400	231,300	0.0	0.0
	7	97,400	90,700	16.9	17.5
	14	67,700	69,800	18.7	22.8
	21	45,800	33,100	26.9	23.2
15	28	17,700	15,900	26.9	23.2

The results show the development of crystallinity in an initially amorphous article during the hydrolysis test.

It will be understood that even though these numerous characteristics and advantages of the invention have been set forth in the foregoing description, together with details of the structure and function of the invention, the disclosure is illustrative only, and changes may be made in detail, especially in matters of shape, size and arrangement of the parts or in the sequence or the timing of the steps, within the broad principle of the present invention to the full extent indicated by the broad general meaning of the terms in which the appended claims are expressed.

WHAT IS CLAIMED:

1. A poly(lactide) film formed from a melt stable polymer composition; said polymer composition comprising:
 - 5 (a) poly(lactide) polymer chains having a number average molecular weight of at least 10,000; said poly(lactide) polymer chains comprising the result of polymerizing a lactide mixture;
 - (b) said polymer composition including a
10 residual lactide concentration of less than about 2% by weight.
2. A film according to claim 1 wherein said polymer composition includes no more than about 2000
15 parts per million water.
3. A film according to claim 1 wherein:
 - (a) said polymer chains comprise the product of polymerizing a lactide mixture comprising greater
20 than about 1% by weight meso-lactide, with any lactide component of said lactide mixture that does not comprise meso-lactide comprising lactide selected from the group consisting essentially of: L-lactide, D-lactide and mixtures thereof;
 - 25 (b) said polymer chains have a number average molecular weight of about 10,000 to about 300,000; and,
 - (c) said film is an amorphous film exhibiting a net endotherm of less than about 10 joules per
30 gram of poly(lactide) polymer.
4. The film of claim 3 wherein said poly(lactide) polymer chains have a number average molecular weight from about 20,000 to about 275,000.

5. The film of claim 3 wherein said poly(lactide) polymer chains have a number average molecular weight from about 40,000 to about 250,000.

5 6. The film of claim 3 wherein said polymer chains are reaction products of polymerizing a lactide mixture comprising about 12 to about 50 percent by weight meso-lactide.

10 7. An article comprising the film of claim 1, said article selected from the group consisting of a diaper, packaging film, agricultural mulch film, bags, shrink wrap film, and tape.

15 8. The film of claim 3 wherein said polymer composition further includes a plasticizer.

9. The film of claim 8 wherein said plasticizer is selected from the group consisting essentially of:
20 alkyl phosphate esters, dialkylether diesters, tricarboxylic esters, epoxidized oils and esters, polymeric polyesters, polyglycol diesters, alkyl alkylether diesters, aliphatic diesters, alkylether monoesters, citrate esters, dicarboxylic esters, esters
25 of glycerine and mixtures thereof.

10. The film of claim 3 wherein said polymer chains are a blend of poly(lactide) and another polymer selected from the group consisting essentially of:
30 cellulose acetate propionate, cellulose butyrate, polycaprolactone and mixtures thereof.

11. The film of claim 10 wherein said film has an elongation of at least 100%.

35

12. A film according to claim 1 wherein:

(a) said film exhibits a net endotherm of less than about 10 joules per gram of poly(lactide) polymer; and,

5 (b) said film comprises a melt stable lactide polymer composition comprising:

(i) poly(lactide) polymer chains, said polymer chains being reaction products of polymerizing a lactide mixture illustrated by the unshaded regions of Figure 4, said polymer chains having a number average molecular weight of from about 10,000 to about 300,000;

(ii) residual lactide in a concentration of less than about 2 percent by weight; and,

15 (iii) less than about 2000 parts per million water.

13. A film according to claim 1 wherein:

(a) said polymer chains comprise the product of polymerizing a lactide mixture comprising less than about 15%, by weight, meso-lactide, with the remaining lactide being selected from the group consisting essentially of: L-lactide, D-lactide and mixtures thereof, wherein at least 85%, by weight, of the lactide in the lactide mixture is either L-lactide or D-lactide;

(b) said polymer chains have a number average molecular weight of from about 10,000 to about 300,000; and,

(c) said film is a semi-crystalline film exhibiting a net endotherm greater than about 10 joules per gram of poly(lactide) polymer.

14. The film of claim 13 wherein said poly(lactide) polymer chains have a number average molecular weight from about 20,000 to about 275,000.

15. The film of claim 13 wherein said poly(lactide) polymer chains have a number average molecular weight from about 40,000 to about 250,000.

5 16. The film of claim 13 wherein said polymer chains are reaction products of polymerizing a lactide mixture comprising less than about 9 percent by weight meso-lactide.

10 17. An article comprising the film of claim 13, said article selected from the group consisting of a diaper, packaging film, agricultural mulch film, shrink wrap film, bags and tape.

15 18. The film of claim 13 wherein said polymer composition further comprises a plasticizer.

19. The film of claim 18 wherein said plasticizer is selected from the group consisting of alkyl phosphate
20 esters, dialkylether diesters, tricarboxylic esters, epoxidized oils and esters, polymeric polyesters, polyglycol diesters, alkyl alkylether diesters, aliphatic diesters, alkylether monoesters, citrate esters, dicarboxylic esters, esters of glycerine and
25 mixtures thereof.

20. A semi-crystalline poly(lactide) film according to claim 13 wherein a substantial portion of crystallinity was induced through orientation.

30

21. The film of claim 20 wherein said orientation was in a plurality of directions.

22. A semi-crystalline film according to claim 13
35 comprising:

(a) 50 to 85% by weight of a semi-crystalline lactide polymer composition exhibiting a net

endotherm greater than about 10 joules per gram of poly(lactide) polymer;

- (b) 15 to 30% by weight of a plasticizer; and,
- (c) 0 to 20% by weight of a filler.

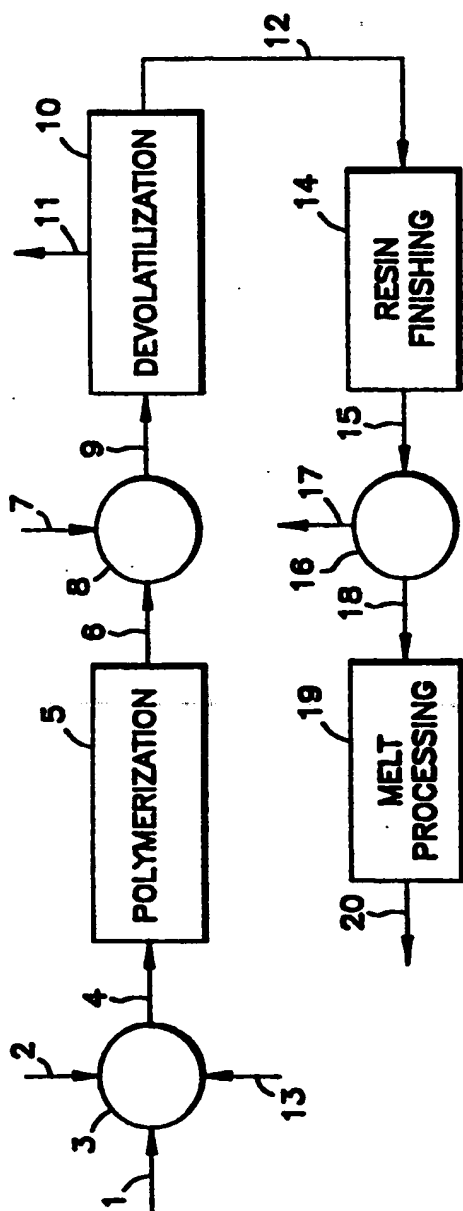


FIG. 1

2/4

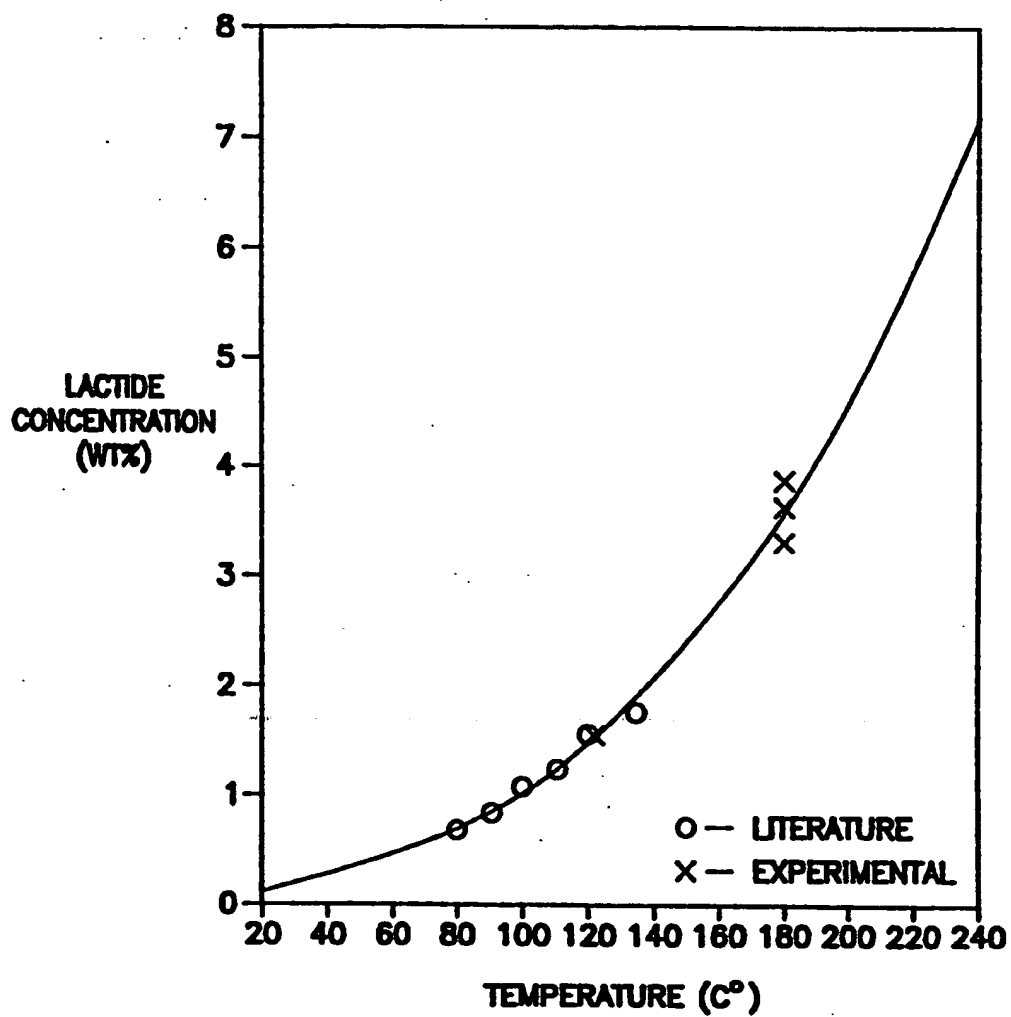
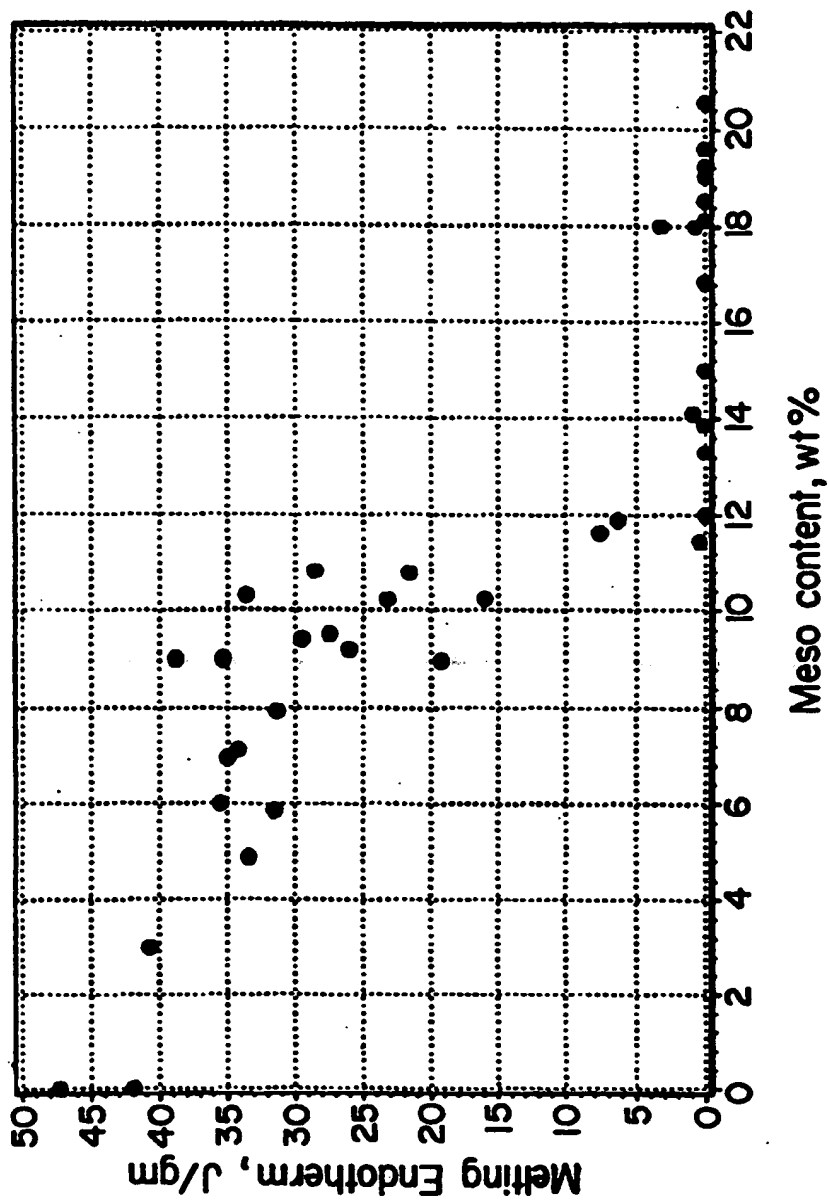


FIG. 2

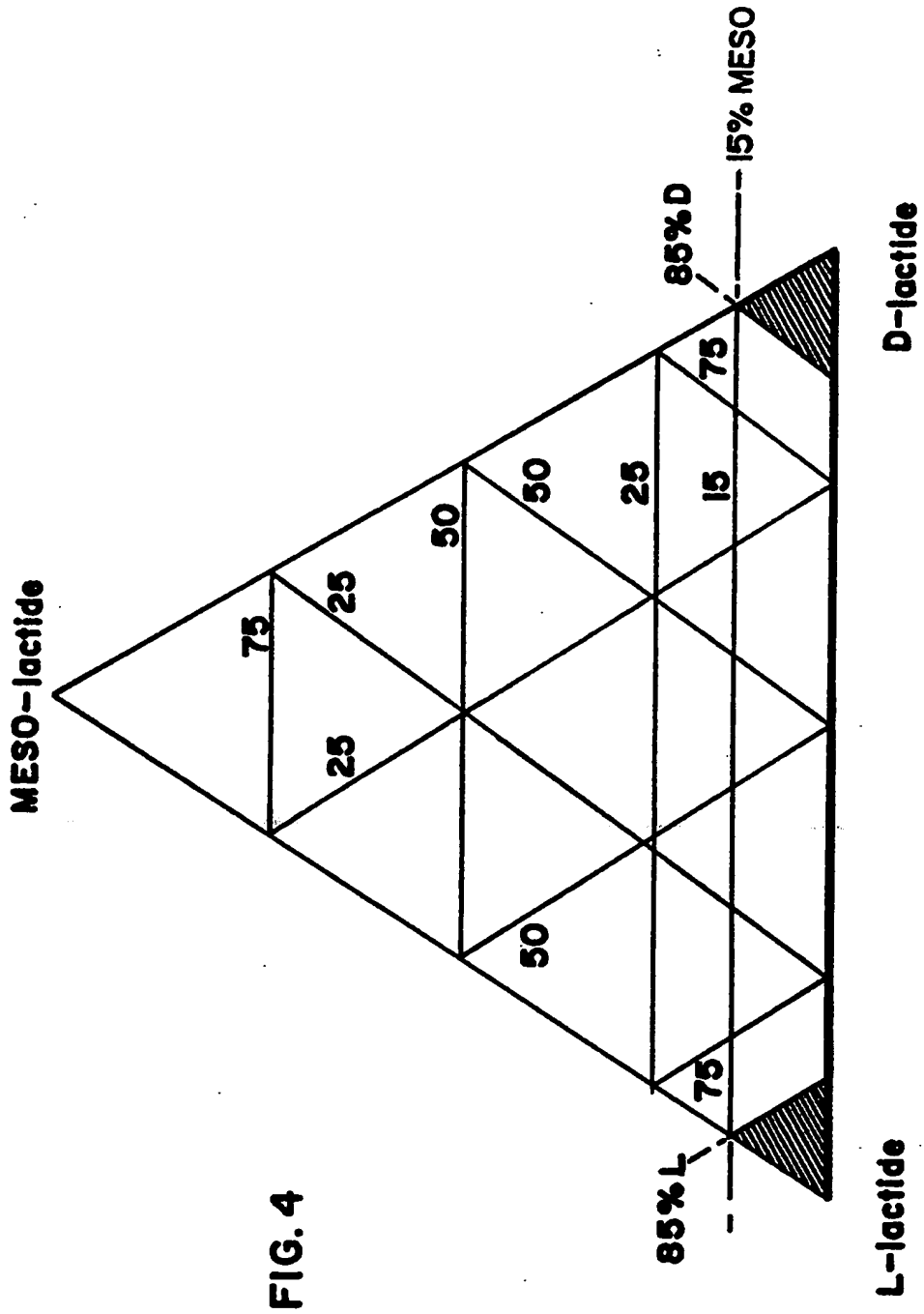
SUBSTITUTE SHEET

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FIG. 3



SUBSTITUTE SHEET



INTERNATIONAL SEARCH REPORT

International Application No
PCT/US 93/09331

A. CLASSIFICATION OF SUBJECT MATTER
IPC 5 C08J5/18

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

IPC 5 C08J C08G

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
P,X	EP,A,0 510 998 (MITSUI TOATSU CHEMICALS) 28 October 1992 see page 7, line 30; table 1 see page 7, line 36 - line 37 see claims 1-10	1
X	WO,A,90 01521 (BATTÉLLE MEMORIAL INSTITUTE) 22 February 1990 see page 34; table 6	1
X	WO,A,92 04412 (DU PONT DE NEMOURS) 19 March 1992 see page 6, line 22 - line 26 see page 12; example 2	1
A	US,A,2 703 316 (ALLAN K. SCHNEIDER) 1 March 1955 * the whole document *	1

☐ Further documents are listed in the continuation of box C.

☒ Patent family members are listed in annex.

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Information on patent family members

International Application No

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